

19 Acids, Bases, and Salts

Big idea REACTIONS

19.1 Acid-Base Theories

For students using the Foundation edition, assign problems 1–7, 9–11, 14, 16–18.

Essential Understanding Acids and bases can be classified in terms of hydrogen ions or hydroxide ions, or in terms of electron pairs.

Reading Strategy

Use Prior Knowledge When you use prior knowledge, you think about your own experience before you read. You will learn new material better if you can relate it to something you already know.

Before you read Lesson 19.1, write your definition for *acid* and *base* in the table below. After you read, write the scientific definition for each term, and compare it to your own definition.

Term	Your definition	Scientific definition
Acid	<i>Answers will vary.</i>	<i>An acid is a compound that produces hydronium ions (H_3O^+) when it is dissolved in water.</i>
Base	<i>Answers will vary.</i>	<i>A base is a compound that produces hydroxide ions (OH^-) when it is dissolved in water.</i>

EXTENSION Name the theory that matches the scientific definition in the table. List three compounds that are acids and three that are bases according to that definition.

Sample answer: Arrhenius; HCl , HNO_3 , H_2SO_4 ; NaOH , KOH , $\text{Ca}(\text{OH})_2$.

Lesson Summary

Arrhenius Acids and Bases Compounds are classified as Arrhenius acids or bases based on whether they ionize to yield hydrogen or hydroxide ions.

- ▶ Arrhenius acids contain a hydrogen atom that is released in water to form a hydronium ion (H_3O^+).
- ▶ Arrhenius bases contain a hydroxide group which is released in water to form a hydroxide ion (OH^-).

Ch19 Acids Bases And Salts Answer Key

Raymond E. Davis, Kenneth W. Whitten

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