

Sample Exercise 24.1 Identifying the Coordination Sphere of a Complex

Palladium(II) tends to form complexes with a coordination number of 4. One such compound was originally formulated as $\text{PdCl}_2 \cdot 3 \text{NH}_3$. **(a)** Suggest the appropriate coordination compound formulation for this compound. **(b)** Suppose an aqueous solution of the compound is treated with excess $\text{AgNO}_3(aq)$. How many moles of $\text{AgCl}(s)$ are formed per mole of $\text{PdCl}_2 \cdot 3 \text{NH}_3$?

Solution

Analyze: We are given the coordination number of Pd(II) and a chemical formula indicating that NH_3 and Cl^- are the potential ligands. We are asked to determine **(a)** what ligands are attached to Pd(II) in the compound and **(b)** how the compound behaves toward AgNO_3 in aqueous solution.

Plan: **(a)** Because of their charge, the Cl^- ions can be either in the coordination sphere, where they are bonded directly to the metal, or outside the coordination sphere, where they are bonded ionically to the complex. Because the NH_3 ligands are neutral, they must be in the coordination sphere. **(b)** The chlorides that are in the coordination sphere will not be precipitated as AgCl .

Solve:

(a) By analogy to the ammonia complexes of cobalt(III), we predict that the three NH_3 groups of $\text{PdCl}_2 \cdot 3 \text{NH}_3$ serve as ligands attached to the Pd(II) ion. The fourth ligand around Pd(II) is one of the chloride ions. The second chloride ion is not a ligand; it serves only as an anion in this ionic compound. We conclude that the correct formulation is $[\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{Cl}$.

(b) The chloride ion that serves as a ligand will not be precipitated as $\text{AgCl}(s)$ following the reaction with $\text{AgNO}_3(aq)$. Thus, only the single “free” Cl^- can react. We therefore expect to produce 1 mol of $\text{AgCl}(s)$ per mole of complex. The balanced equation is $[\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{Cl}(aq) + \text{AgNO}_3(aq) \rightarrow [\text{Pd}(\text{NH}_3)_3\text{Cl}]\text{NO}_3(aq) + \text{AgCl}(s)$. This is a metathesis reaction (Section 4.2) in which one of the cations is the $[\text{Pd}(\text{NH}_3)_3\text{Cl}]^+$ complex ion.

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web een joodse man leest psalmen bij de klaagmuur psalmen hebreeuws סְפָרֵי תְהִלִּים is een van de boeken in de hebreeuwse bijbel de psalmen zijn liederen het bijbelboek bevat volgens de masoretische tekst 150 psalmen een psalter is een apart boek met de inhoud van het bijbelboek en extra s

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alfabetisch register van de psalmen

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web ein psalm ist im judentum und christentum ein poetischer religiöser text oft mit liturgischer funktion die bezeichnung wird vor allem verwendet für die 150 gedichte lieder und gebete des buches der psalmen der hebräischen bibel bzw des alten testaments daneben existieren weitere texte in der biblischen wie in der

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web psalmen 1 1 gelukkig de mens die niet meegaat met wie kwaad doen die de weg van zondaars niet betreedt bij spotters niet aan tafel zit 2 maar vreugde vindt in de wet van de heer en zich verdiept in zijn wet dag en nacht 3 hij zal zijn als een boom geplant aan stromend water

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web apr 10 2021 in einer neuen reihe blickt katholisch de ab heute auf die biblischen psalmen als schule des betens zum auftritt erklärt unser autor die herkunft der biblischen psalmen und ihre bedeutung für das gebet und die persönliche zwiesprache mit gott

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web die zählung der psalmen variiert in den einzelnen traditionen wobei die deutschen bibelausgaben zumeist der zählung der biblia hebraica folgen in der septuaginta wurden psalm 9 10 und 114 115 zusammengefasst so dass sich die

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