

Physical Science- Chapter 24 Test Study Guide KEY

1. _____ - the substances that are about to react.
Reactants
2. _____ - the new substances produced.
Products
3. Law of Conservation of Mass- the French alchemist Antoine Lavoisier established that the mass of the products always _____ the mass of the reactants.
Equals
4. _____ - a way to describe a chemical reaction using chemical formulas and other symbols.
Chemical Equation
5. (s) means? _____
Solid
6. (aq) means? _____
Aqueous, dissolved in water
7.  means? _____
Yields, produces, forms
8. In a chemical reaction atoms can be rearranged but never _____ or _____.
Created, destroyed
9. The numbers to the left of formulas for reactants and products are called _____.
Coefficients
10. _____ - represents the number of units of each substance taking part in a reaction.
Coefficients
11. The following underlined numbers are called _____.
$$\text{NiCl}_2 (\text{aq}) + \underline{2}\text{NaOH} (\text{aq}) \rightarrow \text{Ni}(\text{OH})_2 (\text{s}) + \underline{2}\text{NaCl} (\text{aq})$$

Coefficients
12. Iron exposed to _____ and _____ corrodes or rusts forming iron (III) oxide, rust can seriously damage iron structures because it crumbles and exposes more iron to the air.
Air, moisture
13. Copper corrodes when exposed to the air forming a blue-green coating called a _____.
Patina
14. _____ - has the same number of atoms of each element on both sides of the equation.
Balanced chemical equation

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