

Unit 1: the chemistry of life

Bonds

ionic bonds

- involving a transfer of valence electrons
- usually between a metal and a non-metal
- strongest type of bond
- high melting and boiling points
- high electrical conductivity

covalent bonds

- involving sharing of valence electrons
- usually between two non-metals
- stronger than ionic bonds
- low melting and boiling points
- low electrical conductivity

hydrogen bonds

- weak bonds between polar molecules
- usually between a hydrogen atom and a highly electronegative atom (O, N, F)
- responsible for many of the unique properties of water

Water

Water is a polar molecule due to the difference in electronegativity between oxygen and hydrogen.

Hydrogen bonding in water leads to several unique properties:

- High boiling point for a small molecule
- High surface tension
- Ice floats on liquid water
- Water is a good solvent

Hydrogen bonding

- occurs between polar molecules
- responsible for many of the unique properties of water

Hydrolysis

- breaks down polymers into monomers
- requires water

Dehydration synthesis

- joins monomers into polymers
- releases water

High specific heat

- water can absorb a lot of heat without a large change in temperature

Density

- ice is less dense than liquid water

Universal solvent

- water can dissolve many substances

Carbohydrates

Monomers: $C_6H_{12}O_6$

Polymers: $(C_6H_{10}O_5)_n$

Examples: glucose, fructose, sucrose, starch, glycogen, cellulose

Functions: energy storage, structural support, cell recognition

Carbon

- the basis of organic compounds
- four valence electrons
- can form four covalent bonds

Functional groups

- groups of atoms that determine the chemical and physical properties of a molecule

Examples: hydroxyl, carbonyl, amino, phosphate

Functional Groups

- Hydroxyl** ($-OH$): found in alcohols
- Carbonyl** ($C=O$): found in ketones and aldehydes
- Amino** ($-NH_2$): found in proteins
- Phosphate** ($-PO_4$): found in nucleic acids

Hydrolysis

Dehydration synthesis

Polymers

Long chains of repeating units

Examples: proteins, nucleic acids, polysaccharides

Monomers

Small units that make up polymers

Examples: amino acids, nucleotides, monosaccharides

Macromolecules

- Lipids**: C, H, O
- Proteins**: C, H, O, N
- Nucleic Acids**: C, H, O, N, P

Functions: energy storage, structural support, catalysis, information storage

AP BIO

unit 1 review

From a student who received
A 5 on the 2024 Exam

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Nancy L. Caroline



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