Chapter 8 Guided Reading Enzymes: Biological Catalysts

Complete the following questions as you read Chapter 8 of your textbook. It is suggested that you complete as many questions as you can before attending lecture. After lecture, you can go back through and answer any questions that you weren't able to answer. Questions go mostly in order of the chapter.

Section 8.1 - Enzymes as Biological Catalysts

1. Define the terms: catalyst and substrate.

A catalyst is a substance that increases the rate of a chemical reaction without being consumed in the process. A substrate is a molecule that a catalyst acts upon.

2. What are two very important features of catalysts?

Two very important features of catalysts are: 1) They increase reaction rates without being consumed in the reaction. 2) They do not alter the overall standard Gibbs free energy change of the reaction.

Why is Hb not a true catalyst for the breakdown of H2O2?

Hb is not a true catalyst for the breakdown of H2O2 because Hb gets oxidized in the process. A true catalyst is not consumed in the reaction it catalyzes.

Section 8.2 - The Diversity of Enzyme Function

4. What are the 6 main classes of enzymes? Describe the major function for each class.

The 6 main classes of enzymes arec

- 1. Oxidoreductases: Catalyze oxidation-reduction reactions
- 2. Transferases: Transfer functional groups from one molecule to another
- 3. Hydrolases: Catalyze hydrolysis reactions
- 4. Lyases: Cleave various bonds by means other than hydrolysis and oxidation
- 5. Isomerases: Catalyze isomerization changes within a molecule
- 6. Lizases: Join two molecules together with covalent bonds

Section 8.3 - Chemical Reaction Rates and the Effects of Catalysts

5. Define the term reaction rate.

The reaction rate is the change in concentration of a reactant or product over time. It describes the speed at which a chemical reaction takes place.

6. What is another name for reaction rate?

Another name for reaction rate is velocity.

7. What are the common units of reaction rate?

Common units for reaction rate are molarity per second (M/s) or moles per liter per second (mol/L*s).

8. How do we change the velocity equation to show a decrease in concentration?

To show a decrease in concentration, a negative sign is placed in front of the rate expression.

Ch 8 Guided Reading Chemistry

Julie R. Frentrup

Ch 8 Guided Reading Chemistry:

A Guide to Undergraduate Science Course and Laboratory Improvements National Science Foundation (U.S.).

Directorate for Science Education,1979 Prentice Hall Physical Science Concepts in Action Program Planner

National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

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Peters, 2d Ed Peter P. Berlow,1978 A Guide to Modern Chemistry John F. Baxter, Jay A. Young,1959 Study

Guide/Selected Solutions Manual Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book Student's Guide to Masterton and Slowinski's Chemical Principles Ray Boyington, William L. Masterton, 1977 Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry John McMurry, Susan McMurry, 1986 Study Guide and Solutions Manual to Accompany Fundamentals of Organic Chemistry McMurry, Susan McMurry, 1990 **A Short Guide** to Writing about Chemistry Herbert Beall, John Trimbur, 1996 Advanced advice for students who want to read write and learn about chemistry in preparation for a career in that field **Laboratory Safety Guide** ,2004 **Introductory Chemistry Study Guide** Holder, 2003-05 Books for All Providence Public Library (R.I.), 1926 **Publishers' Weekly** ,1899 The Publishers Weekly, 1902 **Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979 Student Guide for Cycles of Life Gerarld L. Kellogg, 2006 The Chautauguan, 1889 Laboratory Practice ,1977

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