

## Chapter 19 Study Guide: Acids, Bases &amp; Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
Brønsted-Lowry	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

- a.  $\text{H}_2\text{S}$  – hydrosulfuric acid  
 b.  $\text{HC}_2\text{H}_3\text{O}_2$  – acetic acid  
 c.  $\text{NaOH}$  – sodium hydroxide  
 d.  $\text{HNO}_3$  – nitric acid  
 e. Nitrous acid –  $\text{HNO}_2$   
 f. Calcium hydroxide –  $\text{Ca}(\text{OH})_2$   
 g. hydrochloric acid –  $\text{HCl}$

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$M_A \cdot V_A \cdot n_A = M_B \cdot V_B \cdot n_B$$

4. What is the pH of
- $1.34 \times 10^{-4}$
- M solution of hydrochloric acid?

$$\text{pH} = -\log [1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a
- $1.85 \times 10^{-6}$
- M NaOH solution.

$$\text{pOH} = -\log [1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the
- $[\text{OH}^-]$
- of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9} \text{ M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Chap 18 Acid Bases Study Guide Answers

**John A. Olmsted, Gregory M. Williams**



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1. Understanding the eBook Chap 18 Acid Bases Study Guide Answers
  - The Rise of Digital Reading Chap 18 Acid Bases Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chap 18 Acid Bases Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chap 18 Acid Bases Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chap 18 Acid Bases Study Guide Answers
  - Personalized Recommendations

- Chap 18 Acid Bases Study Guide Answers User Reviews and Ratings
- Chap 18 Acid Bases Study Guide Answers and Bestseller Lists
- 5. Accessing Chap 18 Acid Bases Study Guide Answers Free and Paid eBooks
  - Chap 18 Acid Bases Study Guide Answers Public Domain eBooks
  - Chap 18 Acid Bases Study Guide Answers eBook Subscription Services
  - Chap 18 Acid Bases Study Guide Answers Budget-Friendly Options
- 6. Navigating Chap 18 Acid Bases Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chap 18 Acid Bases Study Guide Answers Compatibility with Devices
  - Chap 18 Acid Bases Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chap 18 Acid Bases Study Guide Answers
  - Highlighting and Note-Taking Chap 18 Acid Bases Study Guide Answers
  - Interactive Elements Chap 18 Acid Bases Study Guide Answers
- 8. Staying Engaged with Chap 18 Acid Bases Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chap 18 Acid Bases Study Guide Answers
- 9. Balancing eBooks and Physical Books Chap 18 Acid Bases Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chap 18 Acid Bases Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chap 18 Acid Bases Study Guide Answers
  - Setting Reading Goals Chap 18 Acid Bases Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chap 18 Acid Bases Study Guide Answers
  - Fact-Checking eBook Content of Chap 18 Acid Bases Study Guide Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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