

Physics

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(Chapter 11) (Thermodynamics) (Exercises)

(Class 11)

Question 11.1:

A geyser heats water flowing at the rate of 3.0 litres per minute from 27 °C to 77 °C. If the geyser operates on a gas burner, what is the rate of consumption of the fuel if its heat of combustion is 4.0×10^4 J/g?

Answer 11.1:

Water is flowing at a rate of 3.0 litre/min.

The geyser heats the water, raising the temperature from 27°C to 77°C.

Initial temperature, $T_1 = 27^\circ\text{C}$ and Final temperature, $T_2 = 77^\circ\text{C}$

Rise in temperature, $\Delta T = T_2 - T_1 = 77 - 27 = 50^\circ\text{C}$

Heat of combustion = 4×10^4 J/g

Specific heat of water, $c = 4.2 \text{ J g}^{-1} ^\circ\text{C}^{-1}$

Mass of flowing water, $m = 3.0 \text{ litre/min} = 3000 \text{ g/min}$

Total heat used, $\Delta Q = mc \Delta T = 3000 \times 4.2 \times 50 = 6.3 \times 10^5 \text{ J/min}$

\therefore Rate of consumption = $\frac{6.3 \times 10^5}{4 \times 10^4} = 15.75 \text{ g/min}$

Question 11.2:

What amount of heat must be supplied to 2.0×10^{-2} kg of nitrogen (at room temperature) to raise its temperature by 45 °C at constant pressure? (Molecular mass of $N_2 = 28$; $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$.)

Answer 11.2:

Mass of nitrogen, $m = 2.0 \times 10^{-2} \text{ kg} = 20 \text{ g}$ and Rise in temperature, $\Delta T = 45^\circ\text{C}$

Molecular mass of N_2 , $M = 28$

Universal gas constant, $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$

Number of moles, $n = \frac{m}{M} = \frac{2.0 \times 10^{-2} \times 10^3}{28} = 0.714$

Molar specific heat at constant pressure for nitrogen $C_p = \frac{7}{2}R = \frac{7}{2} \times 8.3 = 29.05 \text{ J mol}^{-1} \text{ K}^{-1}$

The total amount of heat to be supplied is given by the relation: $\Delta Q = nC_p \Delta T$

$= 0.714 \times 29.05 \times 45 = 933.38 \text{ J}$

Therefore, the amount of heat to be supplied is 933.38 J.

Question 11.3:

Explain why

- (a) Two bodies at different temperatures T_1 and T_2 if brought in thermal contact do not necessarily settle to the mean temperature $(T_1 + T_2)/2$.
- (b) The coolant in a chemical or a nuclear plant (i.e., the liquid used to prevent the different parts of a plant from getting too hot) should have high specific heat.
- (c) Air pressure in a car tyre increases during driving.
- (d) The climate of a harbour town is more temperate than that of a town in a desert at the same latitude.

Answer 11.3:

- (a) When two bodies at different temperatures T_1 and T_2 are brought in thermal contact, heat flows from the body at the higher temperature to the body at the lower temperature till equilibrium is achieved, i.e., the temperatures of both the bodies become equal. The equilibrium temperature is equal to the mean temperature $(T_1 + T_2)/2$ only when the thermal capacities of both the bodies are equal.

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Tobias Bleicker



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1. Understanding the eBook Chapter 11 For Physics Workbook
 - The Rise of Digital Reading Chapter 11 For Physics Workbook
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 For Physics Workbook
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 11 For Physics Workbook
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 For Physics Workbook
 - Personalized Recommendations
 - Chapter 11 For Physics Workbook User Reviews and Ratings
 - Chapter 11 For Physics Workbook and Bestseller Lists
5. Accessing Chapter 11 For Physics Workbook Free and Paid eBooks
 - Chapter 11 For Physics Workbook Public Domain eBooks
 - Chapter 11 For Physics Workbook eBook Subscription Services
 - Chapter 11 For Physics Workbook Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Chapter 11 For Physics Workbook Compatibility with Devices
 - Chapter 11 For Physics Workbook Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 11 For Physics Workbook
 - Highlighting and Note-Taking Chapter 11 For Physics Workbook
 - Interactive Elements Chapter 11 For Physics Workbook
8. Staying Engaged with Chapter 11 For Physics Workbook
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 11 For Physics Workbook
9. Balancing eBooks and Physical Books Chapter 11 For Physics Workbook
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 11 For Physics Workbook
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 11 For Physics Workbook
 - Setting Reading Goals Chapter 11 For Physics Workbook
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 11 For Physics Workbook
 - Fact-Checking eBook Content of Chapter 11 For Physics Workbook
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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