

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|-------------------------|---|
| _____ mass | _____ percent composition |
| _____ empirical formula | _____ molar mass |
| _____ molecular formula | _____ standard temperature and pressure (STP, 1 atm, 0°C) |

1. _____ is a description of the relative amounts of each element in a compound.
2. _____ is the lowest whole number ratio of the atoms of the elements in a compound.
3. _____ is the ratio of the mass of one mole of any gas at STP _____ to 22.4 L.
4. _____ is the lowest possible molar mass of a compound.
5. _____ is the ratio of the mass of one mole of a substance to the mass of one mole of ^{12}C .
6. _____ is the ratio of the mass of one mole of a substance to the mass of one mole of ^{12}C .
7. _____ is the ratio of the mass of one mole of a substance to the mass of one mole of ^{12}C .

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following statement by choosing the correct word or phrase from the list below.

- _____ is the amount of substance that contains as many particles as there are atoms in 12 g of carbon-12.
- _____ is the ratio of the mass of one mole of a substance to the mass of one mole of ^{12}C .
- _____ is the ratio of the mass of one mole of a substance to the mass of one mole of ^{12}C .

True or False?

- _____ 1. The molar mass of a substance is the mass of one mole of that substance.
- _____ 2. The molar mass of a substance is the mass of one mole of that substance.
- _____ 3. The molar mass of a substance is the mass of one mole of that substance.
- _____ 4. The molar mass of a substance is the mass of one mole of that substance.

Problems

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