

**10****CHEMICAL QUANTITIES****Chapter Test A****A. Matching**

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

**Column A**

- \_\_\_\_\_ 1. standard temperature and pressure
- \_\_\_\_\_ 2. molar volume
- \_\_\_\_\_ 3. representative particle
- \_\_\_\_\_ 4. molar mass
- \_\_\_\_\_ 5. atomic mass
- \_\_\_\_\_ 6. Avogadro's number
- \_\_\_\_\_ 7. mole (mol)
- \_\_\_\_\_ 8. empirical formula
- \_\_\_\_\_ 9. percent composition
- \_\_\_\_\_ 10. molar mass

**Column B**

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i.  $6.02 \times 10^{23}$  representative particles of a substance
- j.  $0^{\circ}\text{C}$  and 101.3 kPa

**B. Multiple Choice**

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol,  $\text{C}_3\text{H}_8\text{O}$ ?
  - a.  $4 \times (6.02 \times 10^{23})$
  - b. 4
  - c. 32
  - d.  $32 \times (6.02 \times 10^{23})$
- \_\_\_\_\_ 12. The mass of a mole of NaCl is its:
  - a. molar mass.
  - b. empirical formula.
  - c. gram atomic mass.
  - d. atomic mass.

# Chapter 10 Chemical Quantities Test Answers

**I. David Reingold**



## Chapter 10 Chemical Quantities Test Answers:

**Study Guide/Selected Solutions Manual** Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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