

10**CHEMICAL QUANTITIES****Chapter Test A****A. Matching**

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. standard temperature and pressure
- _____ 2. molar volume
- _____ 3. representative particle
- _____ 4. molar mass
- _____ 5. atomic mass
- _____ 6. Avogadro's number
- _____ 7. mole (mol)
- _____ 8. empirical formula
- _____ 9. percent composition
- _____ 10. molar mass

Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i. 6.02×10^{23} representative particles of a substance
- j. 0°C and 101.3 kPa

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol, $\text{C}_3\text{H}_8\text{O}$?
 - a. $4 \times (6.02 \times 10^{23})$
 - b. 4
 - c. 32
 - d. $32 \times (6.02 \times 10^{23})$
- _____ 12. The mass of a mole of NaCl is its:
 - a. molar mass.
 - b. empirical formula.
 - c. gram atomic mass.
 - d. atomic mass.

Chapter 10 Chemical Quantities Workbook Answers

Steven S. Zumdahl



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AQA Key Stage 3 Science Pupil Book 1 Neil Dixon, Carol Davenport, Nick Dixon, Ian Horseywell, 2018-03-19 Ensure that every student develops the maths literacy and working scientifically skills they need to succeed with this skills focused Pupil Book that contains a variety of activities questions and real world examples that are tailored to the Big Ideas and mastery goals of the AQA KS3 Syllabus Develop conceptual understanding with a variety of questions that require students to apply their knowledge to real world scenarios Build working scientifically skills with various Enquiry activities matched to the AQA syllabus Test understanding and measure progress with factual recall questions developed around the ideas of Generalisations Principles and Models Stretch knowledge and understanding with extend tasks linked to higher order thinking skills Compare Evaluate and Predict Bridge the gap between Key Stages 2 and 3 with a focus on maths and enquiry skills and understanding scientific terminology Provides comprehensive support for non specialist or less confident teachers when used in conjunction with the online Teaching Learning resources Written in association with Sheffield Hallam University The Science Education Team within Sheffield Institute of Education SIOE is one of the leading STEM education groups in Europe with a worldwide reputation for knowledge exchange and research SIOE leads national and international STEM education programmes covering curriculum and pedagogical design and development widening participation to traditionally under represented groups and research in science education Basic Chemistry Steven S Zumdahl, 1996 **A Textbook of Physical Chemistry** A. S. Negi, 2007 The Second Revised Edition Of The Book Is Intended To Meet The Requirement Of The Students Of Science Engineering And Other Professional Courses At The Undergraduate Level It Has Been Planned Strictly In Line With The Syllabi Of Various Indian Universities Who Have Adopted The New Ten Plus Two Plus Three Pattern Of Education A New Chapter On Macromolecules Has Been Added Thus Making A Total Of 27 Chapters In The Revised Edition Chapters On Chemical Equilibrium Colligative Properties Atomic Structures Chemical Bonding Have Been Thoroughly Reshuffled And Rewritten Chapter 25 Has Been Rearranged And Divided Into Two Chapters Viz Molecular Spectroscopy And Electrical And Magnetic Properties New Sections Have Been Added To Chapters On Gaseous State Colligative Properties Electrolytic Conduction Ionic Equilibria Chemical Kinetics Atomic Structure And Chemical Bonding Other Chapters Have Also Been Modified And Redesigned The Subject Matter Has Been Given In A Logical Simple And Lucid Language The Main Aim Has Been On Self Learning Some More Diagrams And Illustrations Have Been Added In This Edition For Explaining The Basics And The Fundamentals Of The Subject Conventional Problems In The Earlier Edition Have Been Dropped But General And Objective Type Problems Are Retained A Considerable Number Of Worked Out Problems Have Been Included In Most Of The Chapters These Would Expose The Students To Applications Of Various Concepts And Fundamentals Of The Subject The Revised Text Largely Uses SI Units But CGS Units Have Been Retained In Those Cases Where The SI Units Have Not As Yet Been Fully Appreciated We Have Attempted To Present A Revised Text That Effectively

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