

## Study Guide Chemical Reactions

1. Define oxidation number and provide an example.
2. Define binary compound and provide an example.
3. Discuss which types of elements lose electrons and which gain electrons.
4. What are oxidation numbers and how can you use the periodic table to determine them?
5. Name four elements that tend to have more than one oxidation number.
6. Discuss the charges of ions standing alone and when they are combined.
7. What do you do if you have two ions in an equation with different oxidation numbers?
8. What is a coefficient and what does it mean? Provide an example.
9. Balance the following equation:  
$$\text{H}_2\text{O} + \text{CO}_2 \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2$$
10. Name this binary ionic compound:  $\text{N}_2\text{O}$
11. What is meant by oxidation?
12. What is meant by reduction?
13. Describe in detail a redox reaction.
14. Identify which element was oxidized and which was reduced in the equation you balanced above.
15. Distinguish between reactants and products.
16. What are the four steps to balancing an equation?
17. Describe exothermic and endothermic reactions.
18. Draw an exothermic graph and an endothermic graph.
19. What is meant by the activation energy of a reaction?
20. How does a catalyst affect a reaction?
21. Is the catalyst used up?
22. What is a chemical reaction?
23. Name five indicators that a chemical reaction has occurred.
24. Who was Antoine Lavoisier and what was his contribution to chemistry?
25. What is a chemical formula?
26. What does a chemical formula tell you?
27. Discuss chemical names.
28. Write the chemical formula for each of the following:  
(a) four hydrogens and one carbon (methane); (b) carbon dioxide; (c) dihydrogen oxide
29. List all of the noble gases and discuss their unique characteristics.
30. Why do atoms tend to bond with each other?
31. What is a chemical bond?
32. What is an ion?
33. How do ions form?
34. Describe ionic bonding using  $\text{NaCl}$  as your example.
35. What is a covalent bond?
36. What is a polar covalent bond?
37. What is the advantage of a molecule being polar? Use water as an example.
38. What gives water its ability to dissolve substances?
39. How do multiple bonds form? Use  $\text{N}_2$  as your example and discuss the movement of electrons.
40. What types of elements tend to bond with each other? Use the periodic table.
41. What is metallic bonding and how does it work?
42. How do ions form?

# Chapter 10 Chemical Reactions Study Guide For Content Mastery

**AW Chickering**



## Chapter 10 Chemical Reactions Study Guide For Content Mastery:

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