

Chapter 10 test

De Montfort University

Match each item with the correct statement below:

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

1. the number of representative particles of a substance present in 1 mole of that substance
2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
3. the SI unit used to measure amount of substance
4. 0°C and 1 atm
5. the percent by mass of each element in a compound
6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

7. What SI unit is used to measure the number of representative particles in a substance?
- a. kilogram c. kelvin
b. ampere d. mole
8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C_3H_7O ?
- a. $5 \times (6.02 \times 10^{23})$ c. 35
b. 5 d. $35 \times (6.02 \times 10^{23})$
9. All of the following are equal to Avogadro's number EXCEPT _____
- a. the number of atoms of bromine in 1 mol Br_2
b. the number of atoms of gold in 1 mol Au
c. the number of molecules of nitrogen in 1 mol N_2
d. the number of molecules of carbon monoxide in 1 mol CO
10. How many moles of tungsten atoms are in 4.8×10^{25} atoms of tungsten?
- a. 8.0×10^{-2} moles c. 1.3×10^{-4} moles
b. 8.0×10^{-1} moles d. 1.3×10^{-2} moles
11. How many atoms are in 0.075 mol of titanium?
- a. 1.2×10^{23} c. 6.4×10^{23}
b. 2.2×10^{24} d. 4.5×10^{22}
12. How many molecules are in 2.10 mol CO_2 ?
- a. 2.53×10^{24} molecules c. 3.49×10^{-24} molecules
b. 3.79×10^{24} molecules d. 1.26×10^{24} molecules

Chapter 10 Chemistry Test

Louis Kahlenberg



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