

Chapter 11 Acids and Bases

KEY CONCEPTS

- Describe the principal features of acids and bases.
 - A Brønsted–Lowry acid is a proton donor, often symbolized by HA. A Brønsted–Lowry acid must contain one or more hydrogen atoms.
 - A Brønsted–Lowry base is a proton acceptor, often symbolized by B:. To form a bond to a proton, a Brønsted–Lowry base must contain a lone pair of electrons.
- What are the principal features of an acid–base reaction?
 - In a Brønsted–Lowry acid–base reaction, a proton is transferred from the acid (HA) to the base (B:). In this reaction, the acid loses a proton to form its conjugate base (A:⁻) and the base gains a proton to form its conjugate acid (HB⁺).
- How is acid strength related to the direction of equilibrium in an acid–base reaction?
 - A strong acid readily donates a proton, and when dissolved in water, 100% of the acid dissociates into ions, forming H₃O⁺. A strong base readily accepts a proton, and when dissolved in water, 100% of the base dissociates into ions, forming OH⁻.
 - An inverse relationship exists between acid and base strength. A strong acid forms a weak conjugate base, whereas a weak acid forms a strong conjugate base.
 - In an acid–base reaction, the stronger acid reacts with the stronger base to form the weaker acid and the weaker base.
- What is the acid dissociation constant and how is it related to acid strength?
 - For a general acid HA, the acid dissociation constant K_a is defined by the equation:
$$K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$
 - The stronger the acid, the larger the K_a . Equilibrium in an acid–base reaction favors formation of the acid with the smaller K_a value.
- What is the ion–product of water and how is it used to calculate hydronium or hydroxide ion concentration?

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Chemistry: The Easy Way Joseph A. Mascetta, Mark Kernion, 2019-08-06 A self teaching guide for students Chemistry The Easy Way provides easy to follow lessons with comprehensive review and practice This edition features a brand new design and new content structure with illustrations and practice questions An essential resource for High school and college courses Virtual learning Learning pods Homeschooling Chemistry The Easy Way covers Atomic Structure Chemical Formulas Electrochemistry The Basics of Organic Chemistry And more **5 Steps to a 5 AP Chemistry, 2014-2015 Edition** Richard H. Langley, John Moore, 2013-08-02 A PERFECT PLAN for the PERFECT SCORE STEP 1 Set up your study plan with three customized study schedules STEP 2 Determine your readiness with an AP style diagnostic exam STEP 3 Develop the strategies that will give you the edge on test day STEP 4 Review the terms and concepts you need to score high STEP 5 Build your confidence with full length practice exams **5 Steps to a 5 AP Chemistry, 2010-2011 Edition** John Moore, Richard H. Langley, 2010-01-01 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You ll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the Knowledge Step 5 Build Your Confidence Topics include Reactions and Periodicity Stoichiometry Gases Thermodynamics Spectroscopy Light and Electrons Bonding Solids Liquids and Intermolecular Forces Solutions and Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry and Organic Chemistry Also includes AP Chemistry practice exams AP Advanced Placement Program and College Board are registered trademarks of the College Entrance Examination Board which was not involved in the production of and does not endorse this product **5 Steps to a 5 AP Chemistry, 2012-2013 Edition** Richard H. Langley, John Moore, 2011-06-10 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You ll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the

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