

CHEM 1411 – General Chemistry I
Practice Problems, Chapters 1–3

Chapter 1 – Chemistry: The Study of Change

- 1.** Element, compound, homogeneous mixture (solution), or heterogeneous mixture:
- | | | |
|---------------------------------------|-----------------------|---|
| a) orange juice
(freshly-squeezed) | b) brass | c) 0.9% saline (NaCl) solution |
| d) garden soil | e) room air | f) methane gas |
| g) sodium metal | h) N ₂ gas | i) Cu(NO ₃) ₂ crystals |
- 2.** Define (some of these terms are found in Chapters 2 and 3):
- | | |
|-------------------------|--------------------------------|
| a) element | b) compound |
| c) solution | d) physical change |
| e) chemical change | f) law of conservation of mass |
| g) intensive properties | h) extensive properties |
| i) precision | j) accuracy |
| k) isotopes | l) stoichiometry |
| m) the mole | n) Avogadro's number |
| o) limiting reactant | |
- 3.** Convert
- | |
|--|
| a) 5 feet 10 inches to meters |
| b) 55 miles per hour to cm per second |
| c) 5.0 quarts to liters |
| d) 1.00 ft ² to cm ² |
| e) 32.06 amu to g |
| f) 25 g of carbon to moles of carbon |
| g) 1.00 g of gold to number of atoms of gold |
| h) 25 °F to Kelvins |
- 4.** How many significant figures are in the following numbers?
- | | | |
|-------------|--|---------------|
| a) 3.1416 | b) 6.022 × 10 ²³ atoms/mole | c) 0.000104 g |
| d) 6.0 g/mL | e) 1.00794 amu | f) 5000 mi |

Chapter 1 Chemistry Test

Thomas Evangelist



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