

Chapter 2**The Chemistry of Life****Section 2-1 The Nature of Matter** (pages 33-39)

This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

Atoms (page 33)

1. The basic unit of matter is called a(n) _____.
2. Describe the nucleus of an atom. _____

3. Complete the table about subatomic particles.

SUBATOMIC PARTICLES

Particle	Charge	Location in Atom
	Positive	
	Neutral	
	Negative	

4. Why are atoms neutral despite having charged particles? _____

Elements and Isotopes (page 34)

5. What is a chemical element? _____

6. What does an element's atomic number represent? _____

7. Atoms of the same element that differ in the number of neutrons they contain are known as _____.
8. How are isotopes identified? _____
9. Why do all isotopes of an element have the same chemical properties? _____

Chemical Compounds (page 37)

10. What is a chemical compound? _____

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Edward J. Zalisko



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