

- ____ 13. If the spin of one electron in an orbital is clockwise, what is the spin of the other electron in that orbital?
 a. zero c. counterclockwise
 b. clockwise d. both clockwise and counterclockwise
- ____ 14. What types of atomic orbitals are in the third principal energy level?
 a. *s* and *p* only c. *s*, *p*, and *d* only
 b. *p* and *d* only d. *s*, *p*, *d*, and *f*
- ____ 15. What is the next atomic orbital in the series 1*s*, 2*s*, 2*p*, 3*s*, 3*p*?
 a. 2*d* c. 3*f*
 b. 3*d* d. 4*s*
- ____ 16. What is the number of electrons in the outermost energy level of an oxygen atom?
 a. 2 b. 4 c. 6 d. 8
- ____ 17. What is the electron configuration of potassium?
 a. $1s^2 2s^2 2p^2 3s^2 3p^2 4s^1$ c. $1s^2 2s^2 3s^2 3p^6 3d^1$
 b. $1s^2 2s^2 2p^{10} 3s^2 3p^3$ d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- ____ 18. How many unpaired electrons are in a sulfur atom (atomic number 16)?
 a. 0 b. 1 c. 2 d. 3
- ____ 19. Which is the correct electron dot diagram for the element Silicon (Si)?
 A. $\cdot \ddot{\text{Si}} \cdot$ B. $:\ddot{\text{Si}}:$ C. $\ddot{\text{Si}} \cdot$ D. $:\text{Si}:$
 a. c.
 b. d.
- ____ 20. How many valence electrons are present in silicon?
 a. Two b. Four c. Six d. Eight

Completion: Complete each statement.

21. The _____ principle states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins.
22. The arrangement of electrons in an atom is called the _____ of an atom.
23. The element _____ has the following electron configuration: $[\text{Rn}]5f^6 6d^0 7s^2$.
24. The *f*-block elements are also referred to as the lanthanides and _____.
25. Xenon has _____ valence electrons.
26. Aluminum has _____ valence electrons.

Ch 5 Electrons In Atoms Test Answers

Ralph Thomas Myers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Ch 5 Electrons In Atoms Test Answers:

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Table of Contents Ch 5 Electrons In Atoms Test Answers

1. Understanding the eBook Ch 5 Electrons In Atoms Test Answers
 - The Rise of Digital Reading Ch 5 Electrons In Atoms Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 5 Electrons In Atoms Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 5 Electrons In Atoms Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 5 Electrons In Atoms Test Answers
 - Personalized Recommendations
 - Ch 5 Electrons In Atoms Test Answers User Reviews and Ratings
 - Ch 5 Electrons In Atoms Test Answers and Bestseller Lists
5. Accessing Ch 5 Electrons In Atoms Test Answers Free and Paid eBooks
 - Ch 5 Electrons In Atoms Test Answers Public Domain eBooks
 - Ch 5 Electrons In Atoms Test Answers eBook Subscription Services
 - Ch 5 Electrons In Atoms Test Answers Budget-Friendly Options
6. Navigating Ch 5 Electrons In Atoms Test Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Ch 5 Electrons In Atoms Test Answers Compatibility with Devices
 - Ch 5 Electrons In Atoms Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 5 Electrons In Atoms Test Answers
 - Highlighting and Note-Taking Ch 5 Electrons In Atoms Test Answers
 - Interactive Elements Ch 5 Electrons In Atoms Test Answers
 8. Staying Engaged with Ch 5 Electrons In Atoms Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 5 Electrons In Atoms Test Answers
 9. Balancing eBooks and Physical Books Ch 5 Electrons In Atoms Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 5 Electrons In Atoms Test Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ch 5 Electrons In Atoms Test Answers
 - Setting Reading Goals Ch 5 Electrons In Atoms Test Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ch 5 Electrons In Atoms Test Answers
 - Fact-Checking eBook Content of Ch 5 Electrons In Atoms Test Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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