

## Gas Laws Review Sheet

### Definitions:

1. Boyle's Law – states that as the pressure of a gas increases the volume decreases
2. Charles' Law – states that as the temperature of a gas increases the volume also increases
3. Dalton's Law of Partial Pressures – states that the sum of the partial pressures of individual gases is equal to the total pressure in a container

### Formulas:

4. What is the formula for the combined gas law?  $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
5. What is the formula for the ideal gas law?  $PV = nRT$
6. What is the value for R in the ideal gas law? 8.31 L-kPa / K-mol
7. What does R stand for in the ideal gas law? the ideal gas law constant
8. What does n stand for in the ideal gas law? the number of moles

### Short Answer:

9. What happens to the volume of a gas if the temperature is increased? the volume increases
10. What happens to the volume of a gas if the temperature is decreased? the volume decreases
11. What happens to the volume of a gas if the pressure is increased? the volume decreases
12. What happens to the volume of a gas if the pressure is decreased? the volume increases
13. What happens to the pressure of a gas if the temperature is increased? the pressure increases
14. What happens to the pressure of a gas if the temperature is decreased? the pressure decreases

### Calculations:

15. If 3.2 moles of a gas are at a temperature of 300 K and a volume of 83.8 L, what is the pressure of the gas?

$$P = \frac{nRT}{V}$$

$$P = \frac{3.2 \times 8.31 \times 300}{83.8}$$

$$P = 95 \text{ kPa}$$

16. If 2.4 L of gas is at a pressure of 387 kPa, what would the pressure be if the volume was expanded to 29.8 L?

$$P_2 = \frac{P_1V_1}{V_2}$$

$$P_2 = \frac{387 \times 2.4}{29.8}$$

$$P_2 = 31 \text{ kPa}$$

17. Argon occupies a volume of 28.3 L at a pressure of 50 kPa. Find the volume of argon when the pressure is increased to 310 kPa.

$$V_2 = \frac{P_1V_1}{P_2}$$

$$V_2 = \frac{50 \times 28.3}{310}$$

$$V_2 = 4.6 \text{ L}$$

# Challenging Gas Laws Review All Concepts Answers

**Raffaella Di Napoli**



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