10

CHEMICAL QUANTITIES

Chapter Test A

A. Matching

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

12. The mass of a mole of NaCl is its:

b. empirical formula.

a. modar mass.

	Column A		Column B
1.	standard temperature and pressure		the volume occupied by a mole of any gas at STP (22.4 L)
2	molar volume	b.	the mass in amus of an atom of an element
3.	representative particle	(C)	the number of representative particles in a mole of a substance
4.	molar mass	di.	the percent by mass of each element in a compound
5.	atomic mass	·C.	the mass of a mole of a substance
6.	Avogadro's number	E.	the formula that gives the lowest whole-number ratio of the elements in a compound
7.	mole (mol)	85-	an atom, a formula unit, or a molecule
8.	empirical formula	b.	the mass in grams of a mole of any pure substance
9.	percent composition	30.	6.02×10^{23} representative particles of a substance
10.	molar mass	13-	0°C and 101.3 kPa
B. Multip	ole Choice		
Choose the be	est answer and write its letter	ore i	the time.
11.	How many hydrogen atoms C ₁ H ₂ O?	are	in 4 molecules of isopeopyl alcohol,
	a. $4 \times (6.02 \times 10^{23})$		c. 32
	b. 4		d. $32 \times (6.02 \times 10^{23})$

c. gram atomic mass.

d. atomic mass.

Chapter 10 Chemical Quantities Test

John Iredelle Dillard Hinds

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