

# Chapter 10 Chemical Quantities Review

## Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- |                         |   |
|-------------------------|---|
| _____ mass              | _____ percent composition                                 |
| _____ empirical formula | _____ molar mass  |
| _____ molar mass        | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
1. \_\_\_\_\_ is a description of the relative amounts of each element in a compound.
2. \_\_\_\_\_ is the lowest whole-number ratio of the atoms of the elements in a compound.
3. \_\_\_\_\_ is the mass of 1 mole of any gas at STP.
4. \_\_\_\_\_ is the mass of 1 mole of a compound.
5. \_\_\_\_\_ is the mass corresponding to 1 mol of representative particles of a substance.
6. \_\_\_\_\_ is  $6.02 \times 10^{23}$  particles.
7. \_\_\_\_\_ is the mass of a representative particle of a gas mixture at STP.

## Part 2: Section Concept Review

### 10.1 The Mole: A Measurement

**Completion** Use the completion strategy to develop your knowledge of the terms and your understanding of the concepts involved with the terms of the following concepts.

- Measuring whole units of something, or mass, and its volume for a single quantity called the \_\_\_\_\_ mole. \_\_\_\_\_ Representative particles of a substance is \_\_\_\_\_ Avogadro's number.
- It is the mass of \_\_\_\_\_ of a compound, substance, and together the \_\_\_\_\_ percent composition of the substance of the compound. Other two techniques the substance the mass of the \_\_\_\_\_ molar mass \_\_\_\_\_ of the compound. These are \_\_\_\_\_ STP \_\_\_\_\_ representative particles the mass of one substance.

#### True or False

1. \_\_\_\_\_ 1 mole of any substance is 12 g.

2. \_\_\_\_\_ 1 mole of substance contains 12 g of 12 g atoms.

3. \_\_\_\_\_ 1 mole of representative particles of a substance is the substance. **Incorrect**

4. \_\_\_\_\_ 1 mole of 12 g is composed of one atom of carbon and four atoms of oxygen.

5. \_\_\_\_\_ 1 mole mass of nitrogen gas is 14 g. **Incorrect**

#### Problems

# Chapter 10 Chemical Quantities Answer Key

**Howard Stephen Stoker**



## Chapter 10 Chemical Quantities Answer Key:

**Chemistry: Core Concepts, 3rd Edition** Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package

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**Introduction to Chemistry** Michael Mosher,Paul Kelter,2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry Meyer Brothers

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