

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $\text{PCl}_5$ )
  - b. uranium hexafluoride ( $\text{UF}_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $\text{KMnO}_4$
  - b.  $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $\text{O}_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $\text{C}_6\text{H}_{12}\text{O}_6$
  - b.  $\text{NaHCO}_3$
  - c.  $\text{C}_3\text{H}_8$
  - d.  $\text{KNH}_4\text{SO}_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $\text{PbO}$ )
  - b. 0.75 mol hydrogen sulfide ( $\text{H}_2\text{S}$ )
  - c. 0.00100 mol silicon tetrahydride ( $\text{SiH}_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $\text{O}_2$ )
  - e. 2.30 mol ethylene glycol ( $\text{C}_2\text{H}_6\text{O}_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $\text{N}_2\text{O}_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $\text{Ca}_3(\text{PO}_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $\text{NaBr}$ )
  - b. 13.5 g magnesium nitrate [ $\text{Mg}(\text{NO}_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $\text{MgCl}_2$ )
  - d. 0.00100 g monochloromethane ( $\text{CH}_3\text{Cl}$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $\text{C}_3\text{H}_8(\text{OH})_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $\text{CH}_4$ ) gas?
9. How many moles is 135 L of ammonia ( $\text{NH}_3$ ) gas at STP?

# Chapter 10 Chemical Quantities Answers

**Edward I. Peters**



## Chapter 10 Chemical Quantities Answers:

*Chemical Surety* United States. Department of the Army, 1995      **Basic Chemistry** Steven S Zumdahl, 1996      *Basic Chemical Principles* Donada Peters, Edward I. Peters, 1988      **Essentials of General, Organic, and Biochemistry** Denise Guinn, 2019-03-05 Guinn's *Essentials of General Organic and Biochemistry* uses health and medicine as the framework for learning the fundamentals of chemistry in this student centered one semester text The newly revised 3rd edition focuses on core concepts and necessary math skills and features a revamped organization to align with traditional course organization and shorter more condensed chapters Easily digestible content and medical applications help reduce student anxiety and make chemistry meaningful for students preparing for future careers in nursing and other allied health professions Paired with SaplingPlus and an embedded eBook students will be able to focus their study with adaptive quizzing and understand the relevance of chemistry through videos animations and case studies      **Essentials of General, Organic, and Biochemistry Digital Update** Denise Guinn, 2024-01-17 Guinn's *Essentials of General Organic and Biochemistry 3e Digital Update* is the best intro available for teaching the basics of chemistry to students preparing for careers in nursing and other allied health professions      **Chemistry** Hall, 1995-12      **Study Guide/Selected Solutions Manual** Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book      *Study Guide for Introductory Chemistry* Iris Stovall, 1990      *Basic Concepts of Chemistry* Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance      *Quantitative Chemical Analysis, Sixth Edition* Daniel C. Harris, 2003 For instructors who wish to focus on practical industrial or research chemistry Includes case studies applications boxes and spreadsheet applications      *Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry* Iris Stovall, Steven S. Zumdahl, 1996      *Introductory Chemistry* Steven S. Zumdahl, 1993      **Chemistry** John S. Phillips, Cheryl Wistrom, 2000      **Chemistry for the Biosciences** Jonathan Crowe, Tony Bradshaw, Paul M. S. Monk, 2006 Leading students through the essential concepts that are central to understanding biological systems this text uses everyday examples and analogies to build their confidence in an often daunting subject By focusing on the key themes that unify the subject it shows how integral

chemistry is to the biosciences      Chemistry Zumdahl, Steven S. Zumdahl, 2003-07 Contains chapter discussions key term definitions and practice chapter tests      Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000      **Basic Chemical Principles** Edward I. Peters, 1988      *Essentials of Chemistry* Dennis D. Staley, Antony C. Wilbraham, 1984      Chemical Molecular Science Conrad L. Stanitski, 2004-08      **The Chemical Evolution of Phosphorus** Enrique Macia-Barber, 2019-11-21 Here is a fascinating reader friendly exploration of the phosphorus enigma The volume attempts to answer the questions How did phosphorus atoms which are produced inside the inner cores of a handful of huge stars become concentrated in relatively high proportions in the organisms composing Earth s biosphere And how did these phosphate derivatives manage to be included in such a great variety of organic molecules playing essential biochemical roles in all known life forms Due to the interdisciplinary nature of the topic the volume is arranged in three sections The first section introduces the fundamental concepts and notions of physics chemistry and biology necessary for the proper understanding of the topics discussed within an astronomical framework The author then focuses on the role of phosphorus and its compounds within the context of chemical evolution in galaxies considering its relevance in most essential biochemical functions as well as its peculiar chemistry under different physicochemical conditions The third section provides an overall perspective on the role of phosphorus and its compounds in current areas of research of solid state physics materials engineering nanotechnology or medicine

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