

10**CHEMICAL QUANTITIES****Chapter Test A****A. Matching**

Match each description in Column B to the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. standard temperature and pressure
- _____ 2. molar volume
- _____ 3. representative particle
- _____ 4. molar mass
- _____ 5. atomic mass
- _____ 6. Avogadro's number
- _____ 7. mole (mol)
- _____ 8. empirical formula
- _____ 9. percent composition
- _____ 10. molar mass

Column B

- a. the volume occupied by a mole of any gas at STP (22.4 L)
- b. the mass in amu of an atom of an element
- c. the number of representative particles in a mole of a substance
- d. the percent by mass of each element in a compound
- e. the mass of a mole of a substance
- f. the formula that gives the lowest whole-number ratio of the elements in a compound
- g. an atom, a formula unit, or a molecule
- h. the mass in grams of a mole of any pure substance
- i. 6.02×10^{23} representative particles of a substance
- j. 0°C and 101.3 kPa

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 11. How many hydrogen atoms are in 4 molecules of isopropyl alcohol, $\text{C}_3\text{H}_8\text{O}$?
 - a. $4 \times (6.02 \times 10^{23})$
 - b. 4
 - c. 32
 - d. $32 \times (6.02 \times 10^{23})$
- _____ 12. The mass of a mole of NaCl is its:
 - a. molar mass.
 - b. empirical formula.
 - c. gram atomic mass.
 - d. atomic mass.

Chapter 10 Chemical Quantities Test A Answer Key

Samuel Parsons Mulliken



Chapter 10 Chemical Quantities Test A Answer Key:

Soil Chemical Analysis Marion LeRoy Jackson, 2005 **Chemical Surety** United States. Department of the Army, 1995
Code of Federal Regulations ,1994 **A Method for the identification of pure organic compounds by a systematic analytical procedure based on physical properties and chemical reactions ... v. 1, 1911** Samuel Parsons Mulliken, 1904 The Code of Federal Regulations of the United States of America ,1994 The Code of Federal Regulations is the codification of the general and permanent rules published in the Federal Register by the executive departments and agencies of the Federal Government The Chemical World ,1913 Stowagefactor and Dangerous Goods Segregation Klaus Engeler, 2025-03-24 This Book contains stowagefactors from the following Categories a General Cargoes b Cooling Cargoes c Bulk Cargoes d Ore e Sweet Oils f RoRo g Containersizes h IMDG Code Segregation i German English Dictionary with final Categories **Standard Methods of Chemical Analysis** ,1922 **Elementary chemistry** William Jerome Harrison, 1890 *Official Methods of Analysis of AOAC International* ,2005 Qualitative Chemical Analysis from the Standpoint of Solubilities, Ionization and Mass Action John Iredelle Dillard Hinds, 1913 **Manual of Introductory Chemical Practice ...** George Chapman Caldwell, Abram Adams Breneman, 1875 **Manual of Introductory Chemical Practice for the Use of Students in Colleges and Normal and High Schools** George Chapman Caldwell, Abram Adam Breneman, 1883 Principles of Fire Behavior and Combustion with Advantage Access Richard Gann, 2023-05-31 Principles of Fire Behavior and Combustion Fifth Edition with Navigate Advantage Access is the most current and accurate source of fire behavior information available to firefighters and fire science students today Readers will develop a thorough understanding of the chemical and physical properties of flammable materials and fire the combustion process and the latest in suppression and extinguishment **Standard Methods of Chemical Analysis: The elements.-v.2. Special subjects** Wilfred Welday Scott, 1925 **A Laboratory Guide in Qualitative Chemical Analysis** Horace Lemuel Wells, 1898
Photometric Chemical Analysis John Howe Yoe, 1928 Journal of the American Medical Association American Medical Association, 1925 **Practical Physiology** Edward Provan Cathcart, Diarmid Noël Paton, Marcus Seymour Pembrey, 1923 **A Method for the Identification of Pure Organic Compounds by a Systematic Analytical Procedure Based on Physical Properties and Chemical Reactions ...** Samuel Parsons Mulliken, 1911

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