

# 11 Chemical Reactions

## Big idea REACTIONS

### 11.1 Describing Chemical Reactions

For students using the Foundation edition, assign problems 1–17.

**Essential Understanding** Chemical reactions are represented by balanced chemical equations.

#### Lesson Summary

**Introduction to Chemical Equations** A chemical equation uses symbols, and sometimes words, to show the reactants and products of a chemical reaction.

- ▶ A skeleton equation uses chemical formulas to represent reactants and products, but it does not indicate the relative amounts of each.
- ▶ Anything that enters into a reaction, such as heat or a catalyst, but is not a reactant or product is shown above or below the yields arrow in the equation.

**Balancing Chemical Equations** Chemical equations are balanced to show that mass is conserved during chemical reactions.

- ▶ A balanced equation shows the relative amounts of reactants and products, and it contains equal numbers of each type of atom on both sides of the equation.
- ▶ Chemical equations are balanced by using coefficients in front of the chemical formulas for the reactants and the products in a skeleton equation.



#### BUILD Math Skills

**Balancing Equations** All chemical equations must be balanced because of the law of conservation of mass, which states that matter cannot be created or destroyed. So, the number of atoms that you start with at the beginning of the reaction must equal the number of atoms that you end up with.

For example, the reaction  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  follows the law of conservation of mass because you start with 2 magnesium atoms and 2 oxygen atoms and you end up with 2 magnesium atoms and 2 oxygen atoms. You can think of the  $\rightarrow$  as an  $=$  sign.



Turn the page to learn more about balancing equations.

# Chapter 11 Chemical Reactions Section 11.1

## Answers

**L Reisser**



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