

# The Periodic Law

## SECTION 2

Select the correct answer for the following questions for the space provided.

1. \_\_\_\_\_ for the modern periodic table, elements are ordered
  - (a) according to increasing atomic mass
  - (b) according to Mendeleev's original design
  - (c) according to increasing atomic number
  - (d) based on how they were discovered
2. \_\_\_\_\_ Mendeleev stated that certain elements have similar chemical properties
  - (a) elements appeared in regular intervals and elements discovered later completed intervals of his spacing
  - (b) elements
  - (c) atomic numbers
  - (d) mass ratios
3. \_\_\_\_\_ The modern periodic law states that
  - (a) all new elements with the same properties are listed in the same group
  - (b) the physical and chemical properties of elements are functions of atomic number
  - (c) elements exhibit properties of both periods and groups
  - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. \_\_\_\_\_ The discovery of the noble gases changed Mendeleev's periodic table
  - (a) adding a row
  - (b) adding a column
  - (c) adding a new group
  - (d) adding a new element
5. \_\_\_\_\_ The noble gases that occupy of the noble gases is therefore six
  - (a) helium
  - (b) neon
  - (c) argon
  - (d) krypton
6. \_\_\_\_\_ Likewise, the first element in Group 1 has an atomic number of 3. The second element in this group has an atomic number of
  - (a) 4
  - (b) 10
  - (c) 18
  - (d) 20
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
  - (a) 18 How many elements are in this group?
  - (b) 55 How many elements are in this group?
  - (c) 18 How many elements are in this group?
  - (d) 55 How many elements are in this group?

## Chapter Review The Periodic Law Answers Section 3

**New York (State).**



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