# Self-Assessment Exercise 9.1

Write both forward and reverse reactions and describe macroscopic characteristics of each?

i.  $N_{2 (g)} + 3H_{2 (g)} = 2NH_{3 (g)}$ ii.  $H_{2 (g)} + I_{2 (g)} = 2HI_{(g)}$ Solution: i.  $N_{2 (g)} + 3H_{2 (g)} = 2NH_{3 (g)}$ 

Forward reaction:  $N_{2 (g)} + 3H_{2 (g)} \longrightarrow 2NH_{3 (g)}$ 

Reverse reaction:  $2NH_{3(g)}$   $\longrightarrow$   $N_{2(g)} + 3H_{2(g)}$ 

ii.  $H_{2(g)} + I_{2(g)} = 2HI_{(g)}$ Forward reaction:  $H_{2(g)} + I_{2(g)} = 2HI_{(g)}$ 

Reverse reaction:  $2HI_{(g)}$   $\longrightarrow$   $H_{2(g)} + I_{2(g)}$ 

Macroscopic characteristics of forward and reverse reactions:

Forward Reaction	Reverse reactions
ii. Reactants produce products. iii. Rate is fastest in the beginning	It is written from right to left     ii. Products produce reactants.     iii. Its rate is zero in the beginning and gradually speeds up.

# **Chapter Self Check Answers For Chemistry**

**Steven S Zumdahl** 

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