Covalent Bond Practice

1. Fill in the missing information on the chart.

Element	# of Protons	# of Electrons	# of Valence Electrons	# of electrons to fill outer shell.
Carbon	1	10	4	4
Hydrogen	1	1	/	,
Chlorine	17	17	7	,
Helium	1	2	2	
Phosphorus	15	15	5	3
Oxygen	8	8	4	2
Sulfur	110	110	6	3
Nitrogen	7	7	5	3

- For each of the following covalent bonds:
 - Write the symbols for each element.
 - Draw a Lewis Dot structure for the valence shell of each element.
 - Rearrange the electrons to pair up electrons from each atom.
 - Draw circles to show the sharing of electrons between each pair of atoms.
 - Draw the bond structure using chemical symbols and lines. Use one line for each pair of electrons that is shared.
 - Write the chemical formula for each molecule.

Chapter Six Covalent Bonds Concept Review Answer

Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

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