

CHAPTER 6

STUDY GUIDE

Ionic Compounds and Metals

Lesson 6.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) Chemical Bond. Such an attachment may form by the attraction of the positively charged (2) Nucleus of one atom for the negatively charged (3) Electron of another atom, or by the attraction of charged atoms, which are called (4) Ions. The attractions may also involve (5) Valence electrons, which are the electrons in the outermost (6) Energy level. The (7) Noble Gas are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) Octet. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) Pseudo-Noble gas formations.

For each statement below, write *true* or *false*.

- False 10. A positively charged ion is called an anion.
- True 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- False 12. Elements tend to react so that they acquire the electron structure of a halogen.
- True 13. A sodium atom tends to lose one electron when it reacts.
- True 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- False 15. A Cl^- ion is an example of a cation.
- True 16. The ending *-ide* is used to designate an anion.
- False 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Chapter Study Guide Ionic Compounds And Metals

Leo J. Malone, Theodore O. Dolter



Chapter Study Guide Ionic Compounds And Metals:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. **Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry** Daniel C. Pantaleo, 1986 **Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry** Susan McMurry, 1992 **Study Guide for Atoms, Molecules, and Life** Michael S. Matta, 1981 *Study Guide* Darrell D. Ebbing, Susan M. Schellble, 1998 *Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry* Iris Stovall, Steven S. Zumdahl, 1996 **eBook: General, Organic and Biological Chemistry 2e** SMITH, 2012-02-16 eBook *General Organic and Biological Chemistry 2e* *Study Guide for Chemistry, Third Edition* [by] Steven S. Zumdahl Paul B. Kelter, 1993 Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 eBook *Chemistry: The Molecular Nature of Matter and Change* *Study Guide to accompany Basic Concepts of Chemistry, 7th Edition* Leo J. Malone, 2003-02-20 Work more effectively and gauge your progress along the way. This Study Guide, designed to accompany Malone's Chemistry 7th Edition, includes chapter summaries, new terms, self-tests, answers to self-tests, and solutions to selected problems. This easy-to-read introduction presents chemistry as a living, relevant science. Chemistry 7th Edition encourages critical thinking.

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