



# Chemistry

## Paper Review and Study Guide

**Periodic Table and Law**

**4-page Review + 39 Question Study Guide**

The image displays three overlapping pages from a chemistry study guide. The top page is titled 'Periodic Table and Law' and contains a review of the periodic table and the periodic law. The middle page is titled 'Periodic Table and Law' and contains a review of the periodic table and the periodic law. The bottom page is titled 'Periodic Table and Law' and contains a review of the periodic table and the periodic law.

**Page 1: Periodic Table and Law**

**Review**

- Which group of elements is known for its noble gases?
- Which group contains elements that form +1 ions?
- Which group contains elements that form +2 ions?
- Which group contains elements that form +3 ions?
- Which group contains elements that form +4 ions?
- Which group contains elements that form +5 ions?
- Which group contains elements that form +6 ions?
- Which group contains elements that form +7 ions?
- Which group contains elements that form +8 ions?
- Which group contains elements that form +9 ions?
- Which group contains elements that form +10 ions?
- Which group contains elements that form +11 ions?
- Which group contains elements that form +12 ions?
- Which group contains elements that form +13 ions?
- Which group contains elements that form +14 ions?
- Which group contains elements that form +15 ions?
- Which group contains elements that form +16 ions?
- Which group contains elements that form +17 ions?
- Which group contains elements that form +18 ions?

**Periodic Law**

The periodic law states that the properties of the elements are a periodic function of their atomic number. This means that elements with similar properties are found at regular intervals in the periodic table.

**Page 2: Periodic Table and Law**

**Development of the Modern Periodic Table**

**Early Periodic Tables:**

- Dmitri Mendeleev (1869):** Mendeleev published the first periodic table. He arranged elements by increasing atomic mass and noticed that certain properties periodically repeated. Mendeleev's table left gaps for elements that had not yet been discovered, predicting their properties.
- Luther Meyer (independently):** Luther Meyer published a similar periodic table around the same time as Mendeleev.

**Modern Periodic Table:**

- Moseley's Discovery (1913):** Henry Moseley discovered the atomic number of each element, leading to the modern definition of the periodic law. The modern periodic table is arranged by increasing atomic number rather than atomic mass.

The periodic table is a powerful tool for understanding the behavior of elements. It's organized by increasing atomic number, with elements arranged in rows called periods and columns called groups. Elements within the same group share similar properties, while elements within the same period share similar properties. The periodic table is a key tool for chemists, allowing them to predict the behavior of elements and understand the relationships between them.

**Page 3: Periodic Table and Law**

**Periodic Law**

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# Chapter Study Guide The Periodic Table And Law Answers

**John William Hill, Dorothy M. Feigl**



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