

6. A reaction occurs between two neutral atoms to become ions as represented by the circles below. Determine which one is the metal and label as M, the nonmetal as N, the cation as C^+ and the anion as A^- . Drag and drop:



7. Label each sphere based on size with either F, Br, or Br^- . Drag and drop:



Explain how you determined your labeling above.

8. Write the full electron configuration for the following ions:

a. Co^{3+} : _____

b. Ti^{4+} : _____

c. Ti^{2+} : _____

d. Cl^- : _____

9. Select 5 of the atoms or ions below that are isoelectronic to one another.

Mg²⁺ Ne O²⁻ Ar F⁻ He Al³⁺ Ca²⁺ P³⁻

10. Put the following isoelectronic series in order from smallest to largest. **Sr²⁺ Br⁻ Kr** _____

11. Put the following isoelectronic series in order from smallest to largest. **Na⁺ Al³⁺ O²⁻** _____

12. Consider the elements S, Cl, & K. Drag and drop.

a. List the elements in order of increasing atomic radius: _____ **S Cl K**

b. List their corresponding ions in order of increasing size: _____ **S²⁻ Cl⁻ K⁺**

c. Is there a difference between the two lists? If so, explain why.

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David Kirk



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