

4 • Chemical Equations and Stoichiometry

SIX STOICHIOMETRY PROBLEMS

General Stoichiometry

13. Several brands of antacid tablets use aluminum hydroxide to neutralize excess acid.



[Molar masses: 78.01 36.46 133.4 18.02]

What quantity of HCl, in grams, can a tablet with 0.750 g of $\text{Al}(\text{OH})_3$ consume? What quantity of water is produced?

1.05 g HCl

0.520 g H₂O

17. The equation for one of the reactions in the process of reducing iron ore to the metal is



[Molar masses: 159.7 28.01 55.85 44.01]

- (a) What is the maximum mass of iron, in grams, that can be obtained from 454 g (1.00 lb) of iron(III) oxide?

31.8 g Fe

- (b) What mass of CO is required to reduce the iron(III) oxide to iron metal?

239 g CO

Limiting Reactants

23. The reaction of methane and water is one way to prepare hydrogen:



[Molar masses: 16.04 18.02 28.01 2.02]

If you begin with 995 g of CH_4 and 2510 g of water, what is the maximum possible yield of H_2 ?

CH_4 is LR.... 376 g H_2 is formed

Chapter Stoichiometry Practice Problems Answers

David E. Goldberg



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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