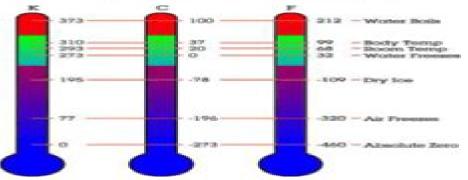
#### Chapter 6 Thermal Energy and Heat

Section 1: Temperature, Thermal Energy, and Heat

- A. Temperature is a measure of the \_\_\_\_\_\_ of the individual particles in a substance.
  - measured with a thermometer and can be measured in \_\_\_\_\_\_\_, Celsius, and

2. \_\_\_\_\_-temperature at which particles stop moving 0°K

K C F 373 — 312 — Worker Bods



3. Temperature conversions

(Fahrenheit) F = 1.8C + 32

(Celsius) C= (F-32)/1.8

4. SI unit for temp. is Kelvin K = C + 273

a. K = C + 273 (10C = 283K)

b. C = K - 273 (10K = -263C)

B. Thermal Energy – the total of all the \_\_\_\_\_\_ of all the particles in a substance.

1. Thermal energy relationships

a. Depends on \_\_\_\_\_\_, mass, and \_\_\_\_\_\_of substance

 b. As temperature \_\_\_\_\_\_, so does thermal energy (because the kinetic energy of the particles increased).

c. Even if the temperature doesn't change, the \_\_\_\_\_\_ in a more massive substance is \_\_\_\_\_\_ (because it is a total measure of energy).

Example: Which beaker of water has more thermal energy?

# **Chapter Thermal Energy And Heat Section Matter Answers**

**Peter Elias Sotiriou** 

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