

CHARLES' LAW

Name _____

Charles' Law states that the volume of a gas varies directly with the Kelvin temperature, assuming that pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$

$$K = ^\circ C + 273$$

Solve the following problems assuming a constant pressure.

1. A sample of nitrogen occupies a volume of ~~250 mL~~ ^{25 L} at 25° C. What volume will it occupy at 95° C?

2. Oxygen gas is at a temperature of 40° C when it occupies a volume of 2.3 liters. To what temperature should it be raised to occupy a volume of 6.5 liters?

3. Hydrogen gas was cooled from 150° C to 50° C. Its new volume is ~~75 mL~~ ^{7.5 L}. What was its original volume?

4. Chlorine gas occupies a volume of ~~25 mL~~ ^{2.5 L} at 300 K. What volume will it occupy at 600 K?

5. A sample of neon gas at 50° C and a volume of 2.5 liters is cooled to 25° C. What is the new volume?

6. Fluorine gas at 300 K occupies a volume of ~~500 mL~~ ^{50 L}. To what temperature should it be lowered to bring the volume to ~~300 mL~~ ^{30 L}?

7. Helium occupies a volume of 3.8 liters at -45° C. What volume will it occupy at 45° C?

8. A sample of argon gas is cooled and its volume went from ~~380 mL~~ ^{38 L} to ~~250 mL~~ ^{25 L}. If its final temperature was -55° C, what was its original temperature?

Charles Law Answer Key With Work

David Kirk



Charles Law Answer Key With Work:

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