

STOICHIOMETRY CHEAT SHEET

Grams (mass) to Moles

$$\frac{\text{grams of known}^*}{1} \times \frac{1 \text{ mol of known}}{\text{grams of known}} = \text{mol of known}$$

Moles to Moles

$$\frac{\text{mol of known}}{1} \times \frac{1 \text{ mol of unknown}}{1 \text{ mol of known}} = \text{mol of unknown}$$

Moles to Grams (mass)

$$\frac{\text{mol of unknown}}{1} \times \frac{\text{grams of unknown}}{1 \text{ mol unknown}} = \text{grams of unknown}^*$$

Grams to Moles to Moles to Grams (Putting it all together)

$$\frac{\text{grams of known}^*}{1} \times \frac{1 \text{ mol of known}}{\text{grams of known}} \times \frac{1 \text{ mol of unknown}}{1 \text{ mol of known}} \times \frac{\text{grams of unknown}}{1 \text{ mol unknown}} = \text{grams of unknown}^*$$

Example

$$\frac{435 \text{ g NaOH}}{1} \times \frac{1 \text{ mol NaOH}}{39.98 \text{ g NaOH}} \times \frac{1 \text{ mol BaCl}_2}{2 \text{ mol NaOH}} \times \frac{208.25 \text{ g BaCl}_2}{1 \text{ mol BaCl}_2} = 1127 \text{ g BaCl}_2$$

Grams of Known

$$\text{Na} = 22.99 \text{ g}$$

$$\text{O} = 15.99 \text{ g}$$

$$\text{H} = 1.00 \text{ g}$$

$$39.98 \text{ g}$$

Example

$$\frac{435 \text{ g NaOH}}{1} \times \frac{1 \text{ mol NaOH}}{39.98 \text{ g NaOH}} = 10.9 \text{ mol NaOH}$$

Example

$$\frac{10.9 \text{ mol NaOH}}{1} \times \frac{1 \text{ mol BaCl}_2}{2 \text{ mol NaOH}} = 5.45 \text{ mol BaCl}_2$$

Balanced Equation: $2\text{NaOH} + \text{BaCl}_2 \rightarrow 2\text{NaCl} + \text{Ba(OH)}_2$

Grams of Unknown

$$\text{Ba} = 137.3 \text{ g}$$

$$2 \times 70.9 = 141.8 \text{ g}$$

$$208.25 \text{ g}$$

Example

$$\frac{5.45 \text{ mol BaCl}_2}{1} \times \frac{208.25 \text{ g BaCl}_2}{1 \text{ mol BaCl}_2}$$

$$= 1127 \text{ g BaCl}_2$$

*Given in problem

*What you are solving for in problem

Cheat Sheet For Stoichiometry

NARAYAN CHANGDER



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A Stoichiometry Unit David Callaghan, 2004 **Stoichiometry** Sydney Young, 1918 This introduction was

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