

Name _____ Class _____ Date _____

Assessment

Chapter Test**Stoichiometry**

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. For the reaction $\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$, calculate the percentage yield if 500. g of sulfur trioxide react with excess water to produce 575 g of sulfuric acid.
- 82.7%
 - 88.3%
 - 91.2%
 - 93.9%
- _____ 2. In the reaction $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$, what is the mole ratio of aluminum to oxygen?
- 10:6
 - 3:4
 - 2:3
 - 4:3
- _____ 3. Fewer steps are required to solve stoichiometry problems when the reactant is given in
- grams and the product is sought in grams.
 - moles and the product is sought in moles.
 - grams and the product is sought in liters.
 - liters and the product is sought in number of atoms.
- _____ 4. Which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield?
- $\text{actual yield} = \frac{\text{percentage yield}}{\text{theoretical yield}} \times 100$
 - $\text{percentage yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
 - $\text{theoretical yield} = \frac{\text{actual yield}}{\text{percentage yield}} \times 100$
- i
 - ii
 - iii
 - Both (b) and (c)

Chapter Stoichiometry Test

John T. Moore, Richard H. Langley



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