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## Chemistry – Study Guide

Ch 10 Study Guide Due: \_\_\_\_\_

Chapter 10 Test – \_\_\_\_\_

When studying for this test, use your 'do nows', notes, homework, class handouts, and your textbook.

### Vocabulary

Avogadro's number

Mole

molar mass

equivalence statement

conversion factor

*For all questions involving calculations, you must show all work, round your answers to the correct number of significant figures, and label your final answers with the correct units.*

### Measuring Matter - Chapter 10 – Section 10.1 – Pgs. 320-324

- Which of the following is NOT a representative particle?
  - Atom
  - Anion
  - Cation
  - All are representative particles
- Avogadro's number of representative particles is equal to one
  - Kilogram
  - Gram
  - Mole
  - Amu
- How many atoms of potassium does 1 mol of potassium contain?

$$6.02 \times 10^{23} \text{ atoms}$$

- How many moles of xenon are equivalent to  $5.66 \times 10^{26}$  atoms?

$$\frac{5.66 \times 10^{26} \text{ atoms}}{1} \times \frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ atoms}} = 940. \text{ mol Xe}$$

# Chapter Study Guide Mass The Mole

**Karen Timberlake**



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