

The Mole Study Guide

Mass Composition: What's your element's % of mass after the reaction? Molar mass to obtain your value!

10. There are 3.75 moles of CO_2 .

11. There are 3.75 moles of CO_2 .

12. There are 3.75 moles of CO_2 .

13. There are 3.75 moles of CO_2 .

14. There are 3.75 moles of CO_2 .

15. There are 3.75 moles of CO_2 .

16. There are 3.75 moles of CO_2 .

17. There are 3.75 moles of CO_2 .

18. There are 3.75 moles of CO_2 .

19. There are 3.75 moles of CO_2 .

20. There are 3.75 moles of CO_2 .

21. There are 3.75 moles of CO_2 .

22. There are 3.75 moles of CO_2 .

23. There are 3.75 moles of CO_2 .

24. There are 3.75 moles of CO_2 .

25. There are 3.75 moles of CO_2 .

26. There are 3.75 moles of CO_2 .

Name _____ Period _____ Date _____

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Fill in the Blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_2\text{H}_4\text{O}_2$	
$\text{C}_3\text{H}_6\text{O}_3$	
$\text{C}_5\text{H}_{10}\text{O}_5$	

Calculate the Molar Mass for the following compounds.

10. H_2O

11. H_2O_2

12. H_2O

13. H_2O_2

Calculate the Percent Composition for each element in the following compound.

10. H_2O

11. H_2O_2

12. H_2O

13. H_2O_2

14. H_2O

15. H_2O_2

16. H_2O

17. H_2O_2

18. H_2O

19. H_2O_2

Determine the Empirical and Molecular Formula for the compound given the information below.

10. A compound has the following percent composition: 58.80% Hg, 21.80% Cl, 1.40% N. The molar mass for Aqueous is 100.0 g/mol.

a. If you have 100 g of Aqueous, how many grams of Hg, Cl, and N do you have?

b. Determine the Empirical Formula for Aqueous.

c. Determine the Molecular Formula for Aqueous.

Chapter Study Guide The Mole

Seager/Slabaugh



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Study guide to accompany Drew H. Wolfe: General, organic and biological chemistry Daniel C. Pantaleo,1986

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Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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