Marme: ( O C)	Date:
Unit 4: Gases Unit Review Sheet	Chemistry R
1. What are the 4 major ideas of Kinetic Molecular Theory?	
a Random, streaght line motion, constan	at
B) elastic collisions	
a volume of paleticles is insignificant	
on no atteactive forces, no ImF	
2. Ideal vs. Real Gases	
PLIGHT He, H2	
atteactive foraces	
nt à small gases	
3. Gas Laws  A) Boyles: $P \downarrow V \uparrow (P \uparrow , V \downarrow)$ Formula: $P \mid V \mid = P \mid V \mid$	)
- Marshmallow/Peep Experiment	
( $\downarrow$ P around the peep, so P inside was now greater than P outside of peep, so V of peep expand)	p 🕇, peep will
B) Charles VT, TT (TJ, VJ)	
Formula:	
➤ Balloon in the freezer example ✓ ✓ · =	
(4 T, the molecules slow down which means fewer collisions with the sides of the ballo pressure inside so V 4)	on, so not mu

# **Chapter1review Gases Answer Key**

**K Morrison** 

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