

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

Chapter Review Chemical Bonding Mixed Answers

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Parmelee**



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Chemical Structure and Bonding
Roger L. DeKock, Harry B. Gray, 1989 Designed for use in inorganic physical and quantum chemistry courses this textbook includes numerous questions and problems at the end of each chapter and an Appendix with answers to most of the problems

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Chemical Engineering for Non-Chemical Engineers Jack Hipple, 2017-01-03 Outlines the concepts of chemical engineering so that non chemical engineers can interface with and understand basic chemical engineering concepts Overviews the difference between laboratory and industrial scale practice of chemistry consequences of mistakes and approaches needed to scale a lab reaction process to an operating scale Covers basics of chemical reaction engineering mass energy and fluid energy balances how economics are scaled and the nature of various types of flow sheets and how they are developed vs time of a project Details the basics of fluid flow and transport how fluid flow is characterized and explains the difference between positive displacement and centrifugal pumps along with their limitations and safety aspects of these differences Reviews the importance and approaches to controlling chemical processes and the safety aspects of controlling chemical processes Reviews the important chemical engineering design aspects of unit operations including distillation absorption and stripping adsorption evaporation and crystallization drying and solids handling polymer manufacture and the basics of tank and agitation system design

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Building Review, 1916 **Paint, Oil and Drug Review**, 1911 **Chemical & Metallurgical**

Engineering Eugene Franz Roeber, Howard Coon Parmelee, 1924

Chemical, Color and Oil Record , 1920

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