

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

1. **A** The notation for sodium chloride, NaCl, stands for one
(a) formula unit. (c) crystal.
(b) molecule. (d) atom.
2. **D** In a crystal of an ionic compound, each cation is surrounded by a number of
(a) molecules. (c) dipoles.
(b) positive ions. (d) negative ions.
3. **B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
(a) higher in potential energy. (c) equal in potential energy.
(b) lower in potential energy. (d) unstable.
4. **B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
(a) Compound A is not an ionic compound.
(b) It will be more difficult to break the bonds in compound A than those in compound B.
(c) Compound B has larger crystals than compound A.
(d) Compound A has larger crystals than compound B.
5. **B** The forces of attraction between molecules in a molecular compound are generally
(a) stronger than the attractive forces among formula units in ionic bonding.
(b) weaker than the attractive forces among formula units in ionic bonding.
(c) approximately equal to the attractive forces among formula units in ionic bonding.
(d) equal to zero.
6. Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

7. What type of energy best represents the strength of an ionic bond?

Lattice energy

8. What types of bonds are present in an ionic compound that contains a polyatomic ion?

Chapter Review Chemical Bonding Section Answers

Aftab S. Hassan, James L. Flowers



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Flowers,1992 This guide for MCAT preparation applies the principles of active and problem based learning to an updated review of content and skills with models for enhanced problem solving and critical thinking abilities There are details on setting up a self managed study programme with guidelines for time management and stress management All areas tested on the exam are covered verbal reasoning physical science writing sample biological sciences with practice questions to chart progress **Chemistry 'O' Level** Rex M. Heyworth,2007 **Chemistry insights 'O' level** Rex M. Heyworth,2007

Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001 Foundations of College

Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

General Chemistry James P. Birk, Donald M. Baird, 1994-03 *Biology* Kenneth Raymond Miller, Prentice Hall (School Division), 1999-02 **Curriculum Review**, 1982 **Prentice Hall Chemistry**, 2000 *Introduction to General, Organic, and Biochemistry* Morris Hein, 1997 The authors recognize that both science and mathematics may be daunting subjects for many students taking this course With this in mind they have anticipated where students might stumble and have paced and organized this text to help them through Their goal is to make the material interesting and relevant so students understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves into organic and biochemistry In this edition the first two sections have been revised primarily to improve explanations and include new pedagogical features The biochemistry portion has been thoroughly updated to include coverage of many recent developments and emerging technologies in the field

The Electrical Review, 1924

Introduction to Organic and Biological Chemistry Michael S. Matta, Antony C. Wilbraham, Dennis D. Staley, 1996

Elements of General and Biological Chemistry John R. Holum, 1987 This book is a shorter version of the third edition of Fundamentals of General Organic and Biological Chemistry 1986 It incorporates the recommendations of the Task Force on Chemical Education for Health Professions and meets the needs for a basic text in a one term course in chemistry for students aiming for careers in professional health care fields

Chemical Interactions, 2005 *Fundamentals of Chemistry* Ralph A. Burns, 2003 For one semester preparatory chemistry courses or general purpose introductory chemistry courses This clearly written well illustrated versatile textbook provides thorough coverage of chemistry with a balance of problem solving skills real world applications and an emphasis on critical thinking and the process of science A supporting theme throughout the text continually emphasizes that chemistry is everywhere

New Zealand Science Review, 1964

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