

The Nature of Matter

READING TOOL Use Structure As you read your textbook, outline the headings and sub-headings throughout this lesson. Fill in the table below with the key points from each section. One has been filled in for you.

| Heading | Sub-Heading | Key Points |
|-----------------------|----------------------|---|
| Atoms | Protons and Neutrons | Inside the atom near the nucleus and have the same #. Protons- positive Neutrons - 0 |
| | Electrons | Electrons are negatively charged and surround the nucleus. |
| Elements and Isotopes | Isotopes | Have a different amount of neutrons in the atom |
| | Radioactive Isotopes | An isotope with an unstable amount of energy making it unstable. Used for research. |
| Chemical Compounds | | A compound with 2 or more elements |
| Chemical Bonds | Ionic Bonds | A charge between 2 opposite charges atoms. |
| | Covalent Bonds | A bond that shares electrons between atoms to make a full shell. |
| | Weak Interactions | Where the energy or elements dont have a strong amount of energy. |

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**Deborah M. Brosnan, Donald J.
Reinhardt**



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