

Stoichiometry Introduction

Stoichiometry is the math side of Chemistry! It relates different quantities of products and reactants in chemical reactions. Stoichiometry relies on balanced equations to find the mole ratios used in most calculations. A mole ratio is the ratio of moles (coefficients) of one reactant/product involved in a reaction to another. We can use mole ratios to convert from moles of one substance to moles of another, but we can also convert between grams and moles.

Example 1: Finding mole ratio



There are 5 moles of O_2 and 2 moles of CO_2 . This means that the mole ratios of O_2 to CO_2 are $\frac{5 \text{ mol } \text{O}_2}{2 \text{ mol } \text{CO}_2}$ and $\frac{2 \text{ mol } \text{CO}_2}{5 \text{ mol } \text{O}_2}$.

1. Find all the mole ratios for the equation $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$.

2. Find all the mole ratios for the equation $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$.

Example 2: mole to mole conversion



If we start with 3 moles of C_2H_6 , how many moles of H_2O will be produced?

$$3 \text{ mol } \text{C}_2\text{H}_6 \times \frac{4 \text{ mol } \text{H}_2\text{O}}{2 \text{ mol } \text{C}_2\text{H}_6}$$

The mol C_2H_6 cancel, leaving you with the mol of H_2O .

$$3 \text{ mol } \text{C}_2\text{H}_6 \times \frac{4 \text{ mol } \text{H}_2\text{O}}{2 \text{ mol } \text{C}_2\text{H}_6} = 12 \text{ mol } \text{H}_2\text{O}$$

Use the following equations for #3-#5: $\text{Ba}^{2+} + 2\text{F}^- \rightarrow \text{BaF}_2$

3. How many moles of KF are produced when 5 mol of Ba^{2+} react?

Chapter Stoichiometry Packet

William S. Seese, Guido H. Daub

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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