

Acids and Bases Worksheet 2

1. In the reaction: $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$
the carbonate ion is acting as a(n):
(A) Arrhenius base
(B) Arrhenius acid
(C) Bronsted-Lowry base
(D) Bronsted-Lowry acid
2. Which of the following represents a Bronsted-Lowry conjugate acid-base pair?
(A) SO_3^{2-} and SO_2 (B) CO_3^{2-} and CO (C) H_3O^+ and H_2 (D) NH_4^+ and NH_3
3. Which of the following would not make a good buffering system?
(A) SO_4^{2-} and H_2SO_4
(B) HCO_3^- and H_2CO_3
(C) NH_3 and NH_4^+
(D) CH_3COO^- and CH_3COOH
4. A water solution of which of the following compounds will turn blue litmus red
(A) K_2CO_3 (B) NH_4Cl (C) NaOH (D) NaCl
5. 10.0 ml of a solution of HCl required 12.5 ml of 0.400 M $\text{Ba}(\text{OH})_2$ for complete neutralization. How many moles of HCl were present in the sample?
(A) 5.00×10^{-3} (B) 1.00×10^{-2} (C) 1.00 (D) 2.00
6. If 1.00 L of 1.00 M CH_3COOH is mixed with 0.25 mole of solid NaOH (assume no volume change), what will be the pH of the resulting solution? ($K_a = 1.8 \times 10^{-5}$).
(A) 4.14 (B) 4.27 (C) 4.74 (D) 5.35
7. What is the molarity of CH_3COOH in vinegar containing 4.0% CH_3COOH by mass and having a density of 1.02 g/ml?
(A) 0.50 M (B) 0.68 M (C) 0.75 M (D) 1.36 M
8. If 5.00 ml of 15.4 M HNO_3 is diluted to 250 ml, what is the pH of the resulting solution?
(A) 0.51 (B) 0.76 (C) 1.45 (D) 2.89

Chapter Supplemental Problems Acids Bases Answer Key

John S. Phillips, Cheryl Wistrom



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Table of Contents Chapter Supplemental Problems Acids Bases Answer Key

1. Understanding the eBook Chapter Supplemental Problems Acids Bases Answer Key
 - The Rise of Digital Reading Chapter Supplemental Problems Acids Bases Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Supplemental Problems Acids Bases Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Supplemental Problems Acids Bases Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Supplemental Problems Acids Bases Answer Key
 - Personalized Recommendations

- Chapter Supplemental Problems Acids Bases Answer Key User Reviews and Ratings
- Chapter Supplemental Problems Acids Bases Answer Key and Bestseller Lists
- 5. Accessing Chapter Supplemental Problems Acids Bases Answer Key Free and Paid eBooks
 - Chapter Supplemental Problems Acids Bases Answer Key Public Domain eBooks
 - Chapter Supplemental Problems Acids Bases Answer Key eBook Subscription Services
 - Chapter Supplemental Problems Acids Bases Answer Key Budget-Friendly Options
- 6. Navigating Chapter Supplemental Problems Acids Bases Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Supplemental Problems Acids Bases Answer Key Compatibility with Devices
 - Chapter Supplemental Problems Acids Bases Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Supplemental Problems Acids Bases Answer Key
 - Highlighting and Note-Taking Chapter Supplemental Problems Acids Bases Answer Key
 - Interactive Elements Chapter Supplemental Problems Acids Bases Answer Key
- 8. Staying Engaged with Chapter Supplemental Problems Acids Bases Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Supplemental Problems Acids Bases Answer Key
- 9. Balancing eBooks and Physical Books Chapter Supplemental Problems Acids Bases Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Supplemental Problems Acids Bases Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Supplemental Problems Acids Bases Answer Key
 - Setting Reading Goals Chapter Supplemental Problems Acids Bases Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Supplemental Problems Acids Bases Answer Key
 - Fact-Checking eBook Content of Chapter Supplemental Problems Acids Bases Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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