

Name: \_\_\_\_\_

Class: \_\_\_\_\_

## OXIDATION NUMBERS WORKSHEET

The *oxidation number* of an atom is the apparent charge assigned to it in a particular molecule, ion or compound. Certain rules are followed in assigning oxidation numbers.

Use rules at the right to assign oxidation numbers to each element in each of the given formulas.

Example

	OXIDATION NO.	RULES USED
$\text{H}_2\text{O}$	<u>H=+1, O= -2</u>	<u>6 &amp; 7</u>
$\text{N}_2$	<u>N= 0</u>	<u>1</u>

**Rules** The oxidation number of:

1. A free element in the uncombined state is 0.
2. A monatomic ion equals the charge on the ion.
3. Oxygen in a compound is generally -2
4. Hydrogen in a compound is +1
5. In a neutral compound the sum of the oxidation numbers for all the atoms is 0.
6. In a polyatomic ion the sum of the oxidation numbers for all the atoms is equals the charge on the ion.
7. Group 1 metals is +1.
8. Group 2 metals is +2.

		RULES USED
1.	$\text{Cl}_2$	_____
2.	$\text{Cl}^-$	_____
3.	$\text{Na}$	_____
4.	$\text{Na}^+$	_____
5.	$\text{KCl}$	_____
6.	$\text{H}_2\text{S}$	_____
7.	$\text{CaO}$	_____
8.	$\text{H}_2\text{SO}_4$	_____
9.	$\text{NO}_3^-$	_____
10.	$\text{Cr}_2\text{O}_7^{2-}$	_____
11.	$\text{NH}_4\text{Cl}$	_____
12.	$\text{NH}_3$	_____
13.	$\text{NO}_2$	_____
14.	$\text{CaH}_2$ (calcium hydride)	_____
15.	$\text{Na}_2\text{O}_2$ (sodium peroxide)	_____

# Charting Oxidation Number Answer Key

**Rosina Ehmman**



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