

Name \_\_\_\_\_

# Stoichiometry

## Questions and Answers

1. What is the mole ratio in the equation  $2H_2 + O_2 \rightarrow 2H_2O$ ?  
2:1:2
2. Define stoichiometry.  
*Stoichiometry is the calculation of reactants and products in chemical reactions based on the balanced equation.*
3. How do you determine the limiting reactant?  
*By calculating the amount of product formed from each reactant and identifying the reactant that forms the least product.*
4. What is the theoretical yield?  
*The maximum amount of product that can be formed from given reactants.*
5. How is percent yield calculated?  
*Percent Yield = (Actual Yield / Theoretical Yield)  $\times$  100%*
6. Given  $N_2 + 3H_2 \rightarrow 2NH_3$ , how many moles of  $NH_3$  are produced from 3 moles of  $N_2$ ?  
6 moles
7. What are excess reactants?  
*Reactants that remain after the limiting reactant is completely used up.*
8. Calculate the theoretical yield if the limiting reactant produces 2 moles of product.  
2 moles (converted to grams if needed)
9. Define actual yield.  
*The amount of product actually obtained from the reaction.*
10. If 5 moles of  $A$  react with 5 moles of  $B$  in the reaction  $A + 2B \rightarrow AB_2$ , which is the limiting reactant?  
 $B$

# Chemistry Study Guide Stoichiometry

**T. R. Dickson**



## **Chemistry Study Guide Stoichiometry:**

**Robinson Chemistry Study Guide** Robinson,1992      **Fundamentals of Chemistry, Study Guide** James E.

Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

**Essential Concepts of Chemistry Study Guide** James R. Braun,Sherman,1999      **Basic Concepts of Chemistry, Study Guide** Leo J. Malone,2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems      **Study Guide for General Chemistry and College Chemistry, Ejghth Editions by Holtzclaw and Robinson** Norman E.

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