

Mole Calculation Practice Worksheet

Answer the following questions:

- 1) How many moles are in 25 grams of water?

$$\frac{25 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{1} = 1.4 \text{ moles}$$

- 2) How many grams are in 4.5 moles of Li_2O ?

$$\frac{4.5 \text{ mol Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \times \frac{29.88 \text{ g Li}_2\text{O}}{1} = 130 \text{ g}$$

- 3) How many molecules are in 23 moles of oxygen?

$$\frac{23 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1} = 1.4 \times 10^{25} \text{ molecules}$$

- 4) How many moles are in 3.4×10^{23} molecules of H_2SO_4 ?

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1} = 0.56 \text{ moles}$$

- 5) How many molecules are in 25 grams of NH_3 ?

$$\frac{25 \text{ g NH}_3}{17.03 \text{ g NH}_3} \times \frac{1 \text{ mol NH}_3}{1} \times \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1 \text{ mol NH}_3} = 8.8 \times 10^{23} \text{ molecules}$$

- 6) How many grams are in 8.2×10^{22} molecules of N_2I_4 ?

$$\frac{8.2 \times 10^{22} \text{ molecules N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \times \frac{1 \text{ mol N}_2\text{I}_4}{1} \times \frac{789.42 \text{ g N}_2\text{I}_4}{1 \text{ mol N}_2\text{I}_4} = 110 \text{ g}$$

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1. Understanding the eBook Chemistry Study Guide The Mole Answers
 - The Rise of Digital Reading Chemistry Study Guide The Mole Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide The Mole Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide The Mole Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide The Mole Answers
 - Personalized Recommendations
 - Chemistry Study Guide The Mole Answers User Reviews and Ratings
 - Chemistry Study Guide The Mole Answers and Bestseller Lists
5. Accessing Chemistry Study Guide The Mole Answers Free and Paid eBooks

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 - Chemistry Study Guide The Mole Answers eBook Subscription Services
 - Chemistry Study Guide The Mole Answers Budget-Friendly Options
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 - Chemistry Study Guide The Mole Answers Compatibility with Devices
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8. Staying Engaged with Chemistry Study Guide The Mole Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide The Mole Answers
9. Balancing eBooks and Physical Books Chemistry Study Guide The Mole Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide The Mole Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide The Mole Answers
- Setting Reading Goals Chemistry Study Guide The Mole Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide The Mole Answers
- Fact-Checking eBook Content of Chemistry Study Guide The Mole Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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