

**University of Mumbai**  
**Revised syllabus S.Y.B.Sc in Chemistry**  
**(w.e.f. the academic year 2009 – 10)**

**Paper I**

Total Lectures 90

Physical and Industrial Analytical Chemistry  
(Numerical problems are expected on all units except Unit 6)

**Term – I**  
**Unit I**

**Chemical Thermodynamics**

1. Gibbs free energy and Helmholtz's free energy, variation of Gibbs free energy with temperature and pressure, Gibbs – Helmholtz equation.
2. Physical equilibria involving pure substances, Clapeyron's equation, and variation of vapour pressure with temperature, Clausius – Clapeyron equation and its application. Partial molal properties, partial molal volume and chemical potential, Gibbs – Duhem equation.
3. Variation of chemical potential with pressure and temperature, fugacity, activity and their relationship with chemical potential, activity and activity coefficient.
4. Thermodynamic derivation of Law of mass action,  $K_p$ ,  $K_c$  and their inter-relation, van't Hoff's reaction isotherm and reaction isochore.

**Unit – II**

**Solution of electrolytes:**

1. Electronic and electrolytic conductors, conductance, specific conductance, equivalent conductance, molar conductance, determination of molar conductance, variation of molar conductance with concentration for strong and weak electrolytes, concept of limiting molar conductance.
2. Debye – Huckel theory of conductance of strong electrolytes, ionic atmosphere relaxation effect, electrophoretic effect.
3. Kohlrausch's law of independent migration of ions.  
Applications of Kohlrausch's law
  - a. Determination of limiting molar conductance of a weak electrolyte,
  - b. Determination of dissociation constant of a weak acid
  - c. Determination of solubility product of a sparingly soluble salt
2. Migration of ions, transport number, dependence of transport number on the velocity of the ion, Hittorf's rule for unattackable electrodes illustrate with ions having equal and the unequal speeds, moving boundary method for the determination of the transport number, factors affecting the transport number of an ion.

# Chemistry Syllabus For Sybsc

**Edward Solly**



## Chemistry Syllabus For Sybsc:

**Syllabus for Chemistry** University of Chicago. Department of Chemistry, 1931 *General Chemistry Syllabus* Orin Fletcher Stafford, 1931 *Inorganic and Organic Chemistry* Dr. Satish M. Chavan, This book is the text book of Inorganic and Organic Chemistry S Y B Sc PAPER II CH 302 Semester III written for second year B Sc students of Savitribai Phule Pune University The book is written according to the New Revised Choice Based Syllabus CBCS of Savitribai Phule Pune University to be implemented from June 2020 This book written in easy and lucid language to understand valence bond theory molecular orbital theory bond formation in molecules co ordination compounds structure and reactivity benzene and their analogs alkyl halides aryl halides alcohols phenols ethers and their nomenclature preparation and reactions For the self study exercise is added with short answer type questions brief answer type questions multiple choice questions MCOs and true false type questions *Chemistry*, 1998 *Chemistry Syllabus* New South Wales. Board of Studies, 1990 **America's First Chemistry Syllabus-and-course for Girls** Wyndham D. Miles, Harold J. Abrahams, 1958 *Chemistry: for B.Sc. Students Semester-I (NEP-UP)* Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc First Semester students of Chemistry as per Common Minimum Syllabus prescribed for all Uttar Pradesh State Universities and Colleges under the recommended National Education Policy 2020 Maintaining the traditional approach to the subject this textbook comprehensively covers two papers namely Fundamentals of Chemistry and Quantitative Analysis Important theoretical topics such as molecular polarity weak chemical forces simple bonding theories of molecules periodic properties of atoms basics of organic chemistry mechanism of organic reactions stereochemistry and the mathematical concepts of chemistry are aptly discussed to give an overview of Fundamentals of Chemistry Practical part covering Quantitative Analysis has been presented systematically to help students achieve solid conceptual understanding and learn experimental procedures **Practical Chemistry** J. Howard, 1873 *Chemistry* University of the State of New York. Bureau of Secondary Curriculum Development, 1967 **Chemistry Syllabus** Pearson Custom Publishing, 1994-05-01 *Chemistry for Degree Students B.Sc. (Honours) Semester II, 1/e (As per CBCS)* Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc Honours Second Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Organic and Physical Chemistry Important topics such as alkanes alkenes alkynes stereochemistry aliphatic hydrocarbons thermochemistry chemical thermodynamics and chemical equilibrium are aptly discussed to give an overview of organic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures *Chemistry for Degree Students B.Sc. Semester - IV (As per CBCS)* Madan R.L., 2017 This textbook has been designed to meet the needs of B Sc Fourth Semester students of Chemistry as per the UGC Choice Based Credit System CBCS With its traditional approach to the subject this textbook lucidly explains

principles of chemistry Important topics such as transition elements coordination chemistry crystal field theory kinetic theory of gases liquids solids and chemical kinetics are aptly discussed to give an overview of inorganic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures

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