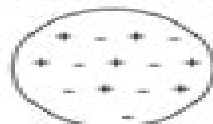


## 200 Ways to Pass the Chemistry Physical Setting Regents Exam

1. **Protons** are positively charged (+).
2. **Neutrons** have no charge.
3. **Electrons** are small and are negatively charged (-).
4. Protons & neutrons are in an atom's nucleus (**nucleons**).
5. Electrons are found in "clouds" (**orbitals**) around an atom's nucleus.
6. The **mass number** is equal to an atom's number of protons and neutrons added together.
7. The **atomic number** is equal to the number of protons in the nucleus of an atom.
8. The **number of neutrons** = mass number – atomic number.
9. **Isotopes** are atoms with equal numbers of protons, but differ in their neutron numbers.
10. **Cations** are positive (+) ions and form when a neutral atom *loses* electrons. They are *smaller* than their parent atom.
11. **Anions** are negative ions and form when a neutral atom *gains* electrons. They are *larger* than their parent atom.
12. **Ernest Rutherford's gold foil experiment** showed that an atom is mostly empty space with a small, dense, positively-charged nucleus.
13. **J.J. Thompson** discovered the electron and developed the "plum-pudding" model of the atom.



Positive & negative  
particles spread throughout  
entire atom.

14. **Dalton's** model of the atom was a solid sphere of matter that was uniform throughout.
15. The **Bohr Model** of the atom placed electrons in "planet-like" orbits around the nucleus of an atom.
16. The current, **wave-mechanical model** of the atom has electrons in "clouds" (orbitals) around the nucleus.
17. USE THE REFERENCE TABLES!!!
18. **"STP"** means "Standard **T**emperature and **P**ressure." (273 Kelvin & 1 atm)
19. Electrons emit energy as light when they jump from higher energy levels back down to lower (**ground state**) energy levels. **Bright line spectra** are produced.
20. **Elements** are pure substances composed of only one kind of atom.
21. **Binary compounds** are substances made up of only two kinds of atoms.  
(examples:  $\text{H}_2\text{O}$ ,  $\text{NH}_3$ ,  $\text{CO}_2$ )
22. **Diatomic molecules** are elements that form two atom molecules in their natural form at STP. Remember the phrase – "**BrINClHO**F" ( $\text{Br}_2$ ,  $\text{I}_2$ ,  $\text{N}_2$ ,  $\text{Cl}_2$ ,  $\text{H}_2$ ,  $\text{O}_2$ ,  $\text{F}_2$ )
23. Use this diagram to help determine the **number of significant figures** in a measured value...

# Chemistry The Physical Setting Topic 2 Answers

**Xiaolong Qi**



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