

Thermodynamics Practice Test

Name: _____

Note - this practice test does NOT contain questions from all sections of the chapters
NChO 1999

____ 27. According to this information, and given the fact that for water, $\Delta H_f = -286 \text{ kJ mol}^{-1}$, what is ΔH_f for MgO(s) ?

Reaction	ΔH
$\text{Mg(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$	-467 kJ mol^{-1}
$\text{MgO(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{O(l)}$	-151 kJ mol^{-1}

- (A) -904 kJ mol^{-1}
- (B) -602 kJ mol^{-1}
- (C) -334 kJ mol^{-1}
- (D) -30 kJ mol^{-1}

NChO 1998

____ 22. Carbon reacts with oxygen according to this equation. $2\text{C(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO(g)}$
 $\Delta H = -220 \text{ kJ}$

Which statements are true?

1. The reaction is exothermic.
 2. The combustion of 0.50 mol of carbon produces 55 kJ of heat energy
- (A) 1 only
 - (B) 2 only
 - (C) both 1 and 2
 - (D) neither 1 nor 2

____ 25. A 1.0 g sample of substance A at 100°C is added to 100 ml of H_2O at 25°C . Using separate 100 mL portions of H_2O , the procedure is repeated with substance B and then with substance C. How will the final temperatures of the water compare?

Substance	Specific Heat
A	$0.60 \text{ J g}^{-1} ^\circ\text{C}^{-1}$
B	$0.40 \text{ J g}^{-1} ^\circ\text{C}^{-1}$
C	$0.20 \text{ J g}^{-1} ^\circ\text{C}^{-1}$

- (A) $T_c > T_b > T_a$
- (B) $T_b > T_a > T_c$
- (C) $T_a > T_b > T_c$
- (D) $T_a = T_b = T_c$

____ 26. How many grams of benzene, $\text{C}_6\text{H}_6\text{(l)}$, must be burned in a bomb calorimeter to raise its temperature by 1.5°C ?
Given: The calorimeter constant is $12.59 \text{ kJ } ^\circ\text{C}^{-1}$ and the $\Delta H^\circ_{\text{combustion}}$ for $\text{C}_6\text{H}_6 = -41.9 \text{ kJ g}^{-1}$

- (A) 0.45 g
- (B) 2.8 g
- (C) 3.3 g
- (D) 8.4 g

Chemistry Thermodynamics Practice Questions

Ralph H. Petrucci, Robert K. Wismer



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