

Baulkham Hills Chemistry Trial

1. Consider the equilibrium process: $\text{CH}_3\text{CH}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CH}_3\text{CH}_{3(g)} \quad \Delta H = -136\text{kJmol}^{-1}$
Which of the following changes would cause the magnitude of the equilibrium constant to increase?

- (a) The temperature is decreased.
- (b) The pressure is decreased.
- (c) The concentration of H_2 in the equilibrium mixture is increased.
- (d) The concentration of CH_3CH_3 in the equilibrium mixture is increased.

2. Phosphorus pentachloride gas undergoes an equilibrium with phosphorus trichloride and chlorine gas according to the following equation:



0.2 moles of phosphorus pentachloride was placed in a sealed 2L container so that equilibrium could be achieved. The container was found to hold 0.15 moles of chlorine gas. What is the equilibrium value (K) for this reaction?

- (a) 0.225
- (b) 0.315
- (c) 0.450
- (d) 1.333

3. At a certain temperature, K for the following reaction is 75:



0.3mol of O_3 and 1.5mol of O_2 were introduced into a 5L reaction vessel. Which is the direction of the equilibrium shift and the reason for the shift?

	Direction Favoured	Reason
(a)	Left	$Q > K$
(b)	Left	$Q < K$
(c)	Right	$Q > K$
(d)	Right	$Q < K$

4. Consider this equilibrium reaction for BOTH questions 4 and 5:



An expression for the equilibrium constant is:

- (a) $\frac{[\text{ClF}_3][\text{F}_2]}{[\text{P}_2][\text{Cl}_2]}$
- (b) $\frac{2[\text{F}_2][\text{Cl}_2]}{2[\text{ClF}_3]}$
- (c) $\frac{[\text{F}_2]^3[\text{Cl}_2]}{[\text{ClF}_3]^2}$
- (d) $\frac{2[\text{ClF}_3]}{2[\text{F}_2][\text{Cl}_2]}$

5. For the equilibrium reaction, the temperature is lowered and the amount of ClF_3 then changed by 0.150mol. The changes occurring would be:

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 - The Rise of Digital Reading Chemistry Trial Papers And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Trial Papers And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Trial Papers And Answers
 - User-Friendly Interface
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 - Personalized Recommendations
 - Chemistry Trial Papers And Answers User Reviews and Ratings
 - Chemistry Trial Papers And Answers and Bestseller Lists
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 - Highlighting and Note-Taking Chemistry Trial Papers And Answers
 - Interactive Elements Chemistry Trial Papers And Answers
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 - Following Authors and Publishers Chemistry Trial Papers And Answers
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 - Carving Out Dedicated Reading Time
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 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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