

Unit 4 Practice MC (No Calculators)

1. A 0.10 mol sample of $\text{Na}_2\text{CO}_3(\text{s})$ and a 0.30 mol sample of $\text{NaCl}(\text{s})$ are dissolved in water and diluted to 250. mL. What is the concentration of Na^+ in the solution?

(A) 0.50 M
(B) 0.500 M
(C) 2.0 M
(D) 2.00 M

$$\begin{aligned} 0.10 \text{ mol Na}_2\text{CO}_3 &= 0.20 \text{ mol Na}^+ \\ 0.30 \text{ mol NaCl} &= 0.30 \text{ mol Na}^+ \end{aligned}$$

0.50 mol Na^+ total

$$[\text{Na}^+] = \frac{0.50 \text{ mol}}{0.25 \text{ L}} = \frac{\frac{1}{2}}{\frac{1}{4}} = \frac{1}{2} \times \frac{4}{1} = 2 \text{ M}$$

2. What is the molarity of $\text{K}^+(\text{aq})$ in a solution that contains 26 g K_3N (molar mass 131 g) in 2.0 L of the solution?

(A) 0.10 M
(B) 0.20 M
(C) 0.30 M
(D) 0.60 M

$$\frac{26 \text{ g K}_3\text{N}}{131 \text{ g}} \times \frac{1 \text{ mol K}_3\text{N}}{131 \text{ g}} = \frac{26}{131} \approx \frac{26}{130} \approx \frac{1}{5} \times 2 = \frac{2}{10} \text{ mol K}_3\text{N}$$

$$\frac{2}{10} \text{ mol K}_3\text{N} = 3 \times \frac{2}{10} \text{ mol K}^+ = \frac{6}{10} \text{ mol K}^+$$

$$[\text{K}^+] = \frac{\frac{6}{10} \text{ mol K}^+}{2 \text{ L}} = \frac{0.6}{2} = 0.3 \text{ M}$$

3. In a qualitative analysis for the presence of Pb^{2+} , Fe^{2+} , and Cu^{2+} ions in an aqueous solution, which of the following will allow the separation of Pb^{2+} from the other ions at room temperature?

(A) Adding Na_2S solution
(B) Adding HCl solution
(C) Adding NaOH solution
(D) Adding HNO_3 solution

PbCl_2 insoluble, but
 FeCl_2 & CuCl_2
are soluble

↑
which causes Pb^{2+}
to precipitate?
But not precipitate
 Fe^{2+} or Cu^{2+}

Chemistry Unit 4 Exam

CL Gary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located on the right side of the slide, partially overlapping the light blue bar.

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