

Chemistry – Unit 6 Worksheet 1

We have observed evidence that atoms of different elements are different, the most noticeable being in their predictable characteristic chemical behaviour. Atoms of one element cannot be transformed into atoms of another element, which demonstrates that each element's particles are made of atoms, regardless of the type of atom involved. However, when the elements react, molecules are formed. Molecules are made of atoms of the same element, which are chemically identical. We will now examine the patterns that exist for the atoms in which there is a regularity in order to describe the changes of the atoms during the reaction.

1. Write the formula and draw the particle diagram for each compound. The order of atoms in each compound is given.

Atomic number of element	16 electrons 16 protons	17 electrons 17 protons	17 electrons 17 protons	18 electrons 18 protons
Element	Sulfur	Chlorine	Chlorine	Argon

particle



Diagram

Atomic number of element	16 electrons 16 protons	17 electrons 17 protons	18 electrons 18 protons
Element	Sulfur	Chlorine	Argon

particle



Diagram

Atomic number of element	16 electrons 16 protons	17 electrons 17 protons	18 electrons 18 protons	18 electrons 18 protons
Element	Sulfur	Chlorine	Chlorine	Argon

particle



Diagram

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Chemistry Unit 6 2 Answers Modeling

VM Jensen



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